KONGUNADU ARTS AND SCIENCE COLLEGE
(AUTONOMOUS)
COIMBATORE – 641 029

DEPARTMENT OF CHEMISTRY (UG)

CURRICULUM AND SCHEME OF EXAMINATIONS (CBCS)
(2023 – 2024 and onwards)

KONGUNADU ARTS AND SCIENCE COLLEGE (AUTONOMOUS)
Coimbatore – 641029

DEPARTMENT OF CHEMISTRY

Vision

To provide personal, intellectual and professional growth of the students and to impart an ideal science education with the emphasis on man-making and character building.

Mission

To produce quality and knowledgeable chemistry graduates capable of creating new developments for the society and preparing chemists of highest caliber for global standards, simultaneously imbibing Indian cultural values in the minds of the students.

PROGRAMME OUTCOME (PO)

On the successful completion of the programme, the following are the expected outcomes

PO 1 Understood the basic concepts, fundamental principles, and the scientific theories related to various scientific phenomena and their relevancies in the day today life.

PO 2 Acquired the skills in handling scientific instruments, planning and performing in laboratory experiments.

PO 3 The skills of observations and drawing logical inferences from the scientific experiments

.

PO 4 Been able to think creatively (divergently and convergent) to propose novel ideas in explaining facts and figures or providing new solution to the problems

PO 5 An understanding of professional, ethical, and social responsibilities

PO 6 Determine the scope and perceive unique areas for further study and employability

PO 7 Students will be skilled in problem solving, critical thinking and analytical reasoning as applied to scientific problems.

PO 8 To be exposed to the different process used in industries and their applications

PROGRAMME SPECIFIC OUTCOME (PSO)

PSO 1 Graduates will have a firm foundation in the fundamentals and application of current chemical and scientific theories including those in Analytical, Inorganic, Organic and Physical Chemistry.

PSO 2 Graduates are able to employ critical thinking and scientific inquiry in the performance, design, interpretation and documentation of laboratory experiments.

PSO 3 Graduates are able to grab enormous existing job opportunities at all levels of chemical, medical, food processing, material industries and educational institutions.

PSO 4 The graduates become entrepreneurs to own enterprises based on the national and International market potentials.

PSO 5 Graduates can perform good social responsibility with greater in ethics and conducive use of natural resources.

KONGUNADU ARTS AND SCIENCE COLLEGE (AUTONOMOUS)
COIMBATORE – 641 029

Course Name: **B. Sc., Chemistry**

Curriculum and Scheme of Examination under CBCS

(Applicable to the students admitted during the Academic Year 2023-2024)

Semester	Part	Subject Code	Title of the Paper	Instruction hours/cycle	Exam. Marks			Duration of Exam (hours)	Credits
					CIA	ESE	TOTAL		
I	I	23TML101	Language I@	6	25	75	100	3	3
	II	23ENG101	English -I	6	25	75	100	3	3
	III	23UCH101	Core Paper 1 – Inorganic, organic and Physical chemistry – I	6	25	75	100	3	5
			Core practical – I Inorganic qualitative analysis & Preparation	3	-	-	-	-	-
	III	23UMA1A1/ 23UZO1A1	Allied A Paper I – Mathematics/ Zoology – I	7/5	25/ 20	75/ 55	100/ 75	3	5/4
			Allied Zoology practical	2	-	-	-	-	-
	IV	23EVS101	Environmental Studies **	2	-	50	50	3	2
			Total	**30**	-	-		-	**18/ 17**
II	I	23TML202	Language II@	6	25	75	100	3	3
	II	23ENG202	English –II	6	25	75	100	3	3
	III	23UCH202	Core Paper 2 – Inorganic, Organic and Physical chemistry – II	6	25	75	100	3	5
	III	23UCH2CL	Core practical – I Inorganic qualitative analysis & Preparation	3	40	60	100	3	2
	III	23UM2A2/ 23UZO2A2	Allied A Paper II – Mathematics/ Zoology – II	7/5	25/ 20	75/ 55	100/ 75	3	5/4
	III	23UZO2AL	Allied Zoology practical	2	20	30	50	3	2
	IV	23VED201	Value Education- Moral and Ethics**	2	-	50	50	3	2
			Total	**30**	-	-		-	**20/**

									21
III	I	23TML303	Language III@	6	25	75	100	3	3
	II	23ENG303	English –III	6	25	75	100	3	3
	III	23UCH303	Core Paper 3 - Inorganic, Organic and Physical chemistry – III	4	25	75	100	3	4
			Core practical – II Inorganic Volumetric and Organic qualitative analysis	3	-	-	-	-	-
	III	23UPH3A1	Allied B Paper 1 – Physics I	4	20	55	75	3	4
			Allied Physics practical - I	3	-	-	-	-	-
	IV	23UGC3S1	Skill Based Subject I- Cyber Security*	2	100	-	100	2	3
	IV	23TBT301/23TAT301/23UHR3N1	Basic Tamil* / Advanced Tamil**/ Non-major elective- I**	2	-	75	75	3	2
			Total	**30**	-	-		-	**19**
IV	I	23TML404	Language IV@	6	25	75	100	3	3
	II	23ENG404	English –IV	6	25	75	100	3	3
	III	23UCH404	Core Paper 4 - Inorganic, Organic and Physical chemistry – IV	4	25	75	100	3	4
	III	23UCH4CM	Core practical – II Inorganic Volumetric and Organic qualitative analysis	3	40	60	100	6	3
	III	23UPH4A2	Allied B Paper II – Physics I	4	20	55	75	3	4
	III	23UPH4AL	Allied Physics practical - I	3	20	30	50	3	2
	IV	23UCH4S2	Skill Based Subject II- Water pollution and management	2	25	75	100	3	3
	IV	23TBT402/23TAT402/23UWR4N2	Basic Tamil* / Advanced Tamil**/ Non-major elective- II**	2	-	75	75	3	2
			Total	**30**	-	-		-	**24**
V	III	23UCH505	Core Paper 5 – Spectroscopy and chromatographic techniques	3	25	75	100	3	3
	III	23UCH506	Core Paper 6 – Inorganic Chemistry	4	25	75	100	3	4
	III	23UCH507	Core Paper 7 – Organic	4	25	75	100	3	4

			reaction mechanism						
	III	23UCH508	Core Paper 8 – Physical Chemistry - I	4	25	75	100	3	4
	III	23UCH5E1	Major Elective 1	3	25	75	100	3	5
	III		Core practical – III Gravimetric Analysis	3	-	-	-	-	-
	III		Core practical – IV Physical Chemistry experiments	4	-	-	-	-	-
	III		Core practical – V Application Oriented Practical	3	-	-	-	-	-
	IV		EDC *	2	100	-	100	3	3
	-	23UCH5IT	Internship Training ****	Grade					
			Total	**30**	-	-		-	**23**
VI	III	23UCH609	Core Paper 9 – Solid state and Coordination Chemistry	5	25	75	100	3	4
	III	23UCH610	Core Paper 10 Chemistry of natural products	5	25	75	100	3	4
	III	23UCH611	Core Paper 11 Physical Chemistry - II	5	25	75	100	3	3
	III	23UCH6E2	Major Elective 2	3	25	75	100	3	5
	III	23UCH6CN	Core practical – III Gravimetric Analysis	3	40	60	100	3	3
	III	23UCH6CO	Core practical – IV Physical Chemistry experiments	4	40	60	100	3	4
		23UCH6CP	Core practical – V Application Oriented Practical	3	40	60	100	3	4
	III	23UCH6Z1	Project viva-voce**	-	20	80	100	-	5
	IV	23UCH6S3	Skill Based Subject III - Food Chemistry	2	25	75	100	3	3
			Total	**30**	-	-		-	**35**
	V	**23NCC$ / NSS/YRC /PYE/ECC/RRC/ WEC101#**	Co curricular Activities*	-	50	-	50	-	1
Grand Total				**-**	**-**	**-**	**3800**	**-**	**140**

Note :

CBCS – Choice Based Credit system, CIA– Continuous Internal Assessment, ESE– End of Semester Examinations

$ For those students who opt NCC under Cocurricular activities will be studying the prescribed syllabi of the UGC which will include Theory, Practical & Camp components. Such students who qualify the prescribed requirements will earn an additional 24 credits.

@ Hindi/Malayalam/ French/ Sanskrit – 23HIN/MLM/FRN/SAN101 - 404

***** - No End-of-Semester Examinations. Only Continuous Internal Assessment (CIA)

****-** No Continuous Internal Assessment (CIA). Only End-of-Semester Examinations (ESE)

*** Project Report – 60 marks; Viva voce – 20 marks; Internal-20 marks

**** The students shall undergo Internship training / field work for a minimum period of 14 working days at the end of the fourth semester during summer vacation and submit the report in the fifth semester which will be evaluated for 100 marks by the concerned guide and followed by an Internal Viva voce by the respective faculty or HOD as decided by the department. According to their marks, the grades will be awarded as given below.

Marks %	Grade
85 – 100	O
70 – 84	D
60 – 69	A
50 – 59	B
40 – 49	C
< 40	U (Reappear)

Major Elective Papers

(2 papers are to be chosen from the following 6 papers)

1. Polymer technology
2. Nano and green Chemistry
3. Pharmaceutical Chemistry
4. Agricultural Chemistry
5. Dairy Chemistry
6. Leather Chemistry

Non-Major Elective Papers

1. Human Rights
2. Consumer Affairs
3. Womens Rights

Sub. Code & Title of the Extra Departmental Course (EDC):

23UCH5X1 - Chemistry in day today life

 # List of Co curricular Activities:

1. National Cadet Corps (NCC)
2. National Service Scheme (NSS)
3. Youth Red Cross (YRC)
4. Physical Education (PYE)
5. Eco Club (ECC)
6. Red Ribbon Club (RRC)
7. Women Empowerment Cell (WEC)

Note: In core/ allied subjects, no. of papers both theory and practical are included wherever applicable. However, the total credits and marks for core/allied subjects remain the same as stated below.

Tally Table

S.No.	Part	Subject	Marks	Credits
1.	I	Language – Tamil/Hindi/Malayalam/ French/ Sanskrit	400	12
2.	II	English	400	12
3.	III	Core – Theory/Practical	1600	60
	III	Allied	400	20
		Electives/Project	300	15
4.	IV	Basic Tamil / Advanced Tamil (OR) Non-major electives	150	4
		Skill Based subject	300	9
		EDC	100	3
		Environmental Studies	50	2
		Value Education	50	2
5.	V	Cocurricular Activities	50	1
		Total	**3800**	**140**

- 25 % CIA is applicable to all subjects except JOC, COP and SWAYAM courses which are considered as extra credit courses.
- 100 % CIA for Cyber Security and EDC paper.
- The students to complete any **MOOC On learning platforms like SWAYAM, NPTEL, Course era, IIT Bombay Spoken Tutorial etc.,** before the completion of the 5th semester and the course completion certificate should be submitted through the HOD to the Controller of Examinations. Extra credits will be given to the candidates who have successfully completed.
- An **Onsite Training** preferably relevant to the course may be undertaken as per the discretion of the HOD.

- Students who successfully complete **Naan Mudhalvan** courses in 3rd and 5th semester will be given 2 extra credits for each course. They are asked to submit the marks to Controller of Examinations through and undersigned by the HOD.

Semester	Naan Mudhalvan Course Title
III	Digital skills for employability
V	Advanced technology for employability in life Science

Components of Continuous Internal Assessment

Components		**Marks**	**Total**
Theory			
CIA I	75	(75+75) Converted to 15	25
CIA II	75		
Assignment/Seminar		5	
Attendance		5	
Theory			
CIA I	75	(50+50) Converted to 15	25
CIA II	75		
Assignment/Seminar		5	
Attendance		5	
Theory (Allied)			
CIA I	55	(55+55) Converted to 10	20
CIA II	55		
Assignment/Seminar		5	
Attendance		5	
Practical			
CIA Practical		25	40
Observation Notebook		10	
Attendance		5	
Practical			
CIA Practical		10	25
Observation Notebook		10	
Attendance		5	
Practical (Allied)			
CIA Practical		10	20
Observation Notebook		5	
Attendance		5	
Project/Case study			
Review		15	20
Regularity		5	

BLOOM'S TAXONOMY BASED ASSESSMENT PATTERN

(K1-Remembering; **K2**-Understanding; **K3**-Applying; **K4**-Analyzing; **K5**-Evaluating)

Theory Examination – Part I, II & III

(i) CIA I & II and ESE: 75 Marks

Knowledge Level	Section	Marks	Description	Total
K1 Q1 to 10	A (Answer all)	10 x 1 = 10	MCQ	75
K1 – K5 Q11 to 15	B (Either or pattern)	5 x 5 = 25	Short Answers	
K2 – K5 Q16 to 20	C (Either or pattern)	5 x 8 = 40	Descriptive / Detailed	

(ii) CIA I & II and ESE: 55 Marks (Allied)

Knowledge Level	Section	Marks	Description	Total
K1 Q1 to 10	A (Answer all)	10 x 1 = 10	MCQ	55
K1 – K5 Q11 to 15	B (Either or pattern)	5 x 3 = 15	Short Answers	
K2 – K5 Q16 to 20	C (Either or pattern)	5 x 6 = 30	Descriptive / Detailed	

(iii)CIA I & II and ESE: 50 Marks

Knowledge Level	Section	Marks	Description	Total
K1 Q1 to 10	A (Answer all)	10 x 1 = 10	MCQ	50**
K1 – K5 Q11 to 15	B (Either or pattern)	5 x 3 = 15	Short Answers	
K2 – K5 Q16 to 20	C (Either or pattern)	5 x 5 = 25	Descriptive / Detailed	

****For ESE 50 marks converted to 25 marks.**

2. ESE Practical Examination:

Knowledge Level	Section	Marks	Total
K3	Experiments	50	60
K4	Record Work	10	
K5			

(For Allied papers)			
Knowledge Level	**Section**	**Marks**	**Total**
K3	Experiments	25	30
K4	Record Work	05	
K5			

Knowledge Level	**Section**	**Marks**	**Total**
K3	Experiments	20	25
K4	Record Work	05	
K5			

3. ESE Project Viva Voce:

Knowledge Level	**Section**	**Marks**	**Total**
K3	Project Report	60	80
K4	Viva voce	20	
K5			

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – I: INORGANIC, ORGANIC AND PHYSICAL CHEMISTRY – I**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	I	6	90	5

Objectives

1.	To know the concept of qualitative inorganic analysis.
2.	To acquaint knowledge about electron displacement effects, hybridization and conformations.
3.	To know about the structure of an atom.

Course Outcomes (CO)

K1 – K5	CO1	Explain the basic analytical knowledge and group separation of elements.
	CO2	To know the types of bonding and geometry in molecules and VSEPR theory
	CO3	Explain the isomerism of alkanes and cycloalkanes.
	CO4	Acquire the knowledge about the structure of atoms.
	CO5	Understand characteristics of gases.

UNIT – I	**Qualitative analysis**	**18 Hours**

Introduction - Dry reactions – heating, flame tests; Wet reactions – test tubes, centrifuge tubes, stirring rods, droppers, reagent bottles and reagents, the centrifuge, washing the precipitates, wash bottles, transferring of precipitates, heating of solutions, evaporation, dissolving of precipitates, precipitation with hydrogen sulphide, cleaning of apparatus. Interfering anions and its elimination, classification of cations into analytical groups (group separation only), scheme of classification of anions.

Industry 4.0 Introduction to Industry 4.0- Need – Reasons for Adopting Industry 4.0 - Definition – Goals Technologies of Industry 4.0- Applications of Artificial Intelligence in chemistry for predicting the properties of molecular structure – Chem sketch, Chem Draw, MOPAC, Avagadro.

UNIT-II	**Chemical Bonding**	**18 Hours**

Chemical bond – definition - Types of chemical bonds.Ionic or electrovalent bond: Definition - Illustration of the formation of ionic bond, (Examples: NaCl, MgO, CaF_2, Al_2O_3 only) - Condition for the formation of ionic compounds - Characteristics of ionic compounds - Born Haber cycle.

Covalent bond: Definition - Types of covalent bond (single, double and triple) - Illustration of the formation of covalent bond (Example: HF, H_2O, NH_3, O_2, N_2 only) - Characteristics of covalent compounds.

Coordinate bond: Definition - Illustration of the formation of coordinate bond (Example: SO_2, $NH^{4+,}$ Al_2Cl_6 only) - Comparison between Ionic, Covalent and Coordinate Bond. Hydrogen bond: Definition - Illustration of the formation of Hydrogen bond (Example H_2O, 2 - Nitrophenol only) - Conditions for hydrogen bonding properties - Types of Hydrogen Bonding and Characteristic of Hydrogen Bonded Compounds.

Hybridisation and Geometries of molecules – VSEPR theory - Geometry of $BeCl_2$, PCl_3, H_2O only.		
UNIT-III	**Chemistry of alkanes and Cycloalkanes**	**18 Hours**
Inductive effect, electromeric effect, mesomeric effect and hyperconjucative effect. Homolytic and heterolytic fission, Reaction Intermediates-carbocations, carbanions, carbon free radicals and carbenes. Classification of reagents: Electrophiles and nucleophiles. Types of organic reactions- Substitution, addition, elimination and rearrangement reactions (Basic ideas only). Structure – Nomenclature - Isomerism in alkanes – Natural source of alkanes – Methods of preparation – physical properties - chemical properties – Conformation study of ethane and n-butane. Nomenclature – methods of preparation – physical and chemical properties. Stability of cycloalkanes, Bayer strain theory.		
UNIT-IV	**Structure of atom – Classical Mechanics and Wave mechanical approach**	**18 Hours**
Introduction of classical mechanics approach - Quantum Theory and Bohr Atom. Wave mechanical concept of atom – de Broglie's equation. Heisenberg's Uncertainty principle. Schrodinger's Wave equation. Charge cloud concept and orbitals. Quantum Numbers-Principal, Azimuthal, Magnetic and Spin Quantum Numbers and their significance. Pauli's Exclusion principle. Energy distribution and orbitals. Distribution of electrons in orbitals. Representation of ground state electronic configuration of elements – ***Aufbau principle and its limitations, n+l rule and Hund's rule.** Pauling's exclusion principle for multi electron system**.**		
UNIT-V	**Gaseous state**	**18 Hours**
Characteristics of gases- parameters of a gas. Gas laws-Boyele's law, Charle's Law, Gay lussac's law and Avogadro's law. The ideal gas equation- kinetics of molecular theory of gases. Derivation of kinetic gas equation. Distribution of molecular velocities- calculation of molecular velocities. Collision properties. Van der Waals equation – Liquefaction of gases – Law of corresponding states- Methods of liquefaction of gases.		
***Self study portion**		
Teaching Methods: Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS:		
1.	G. Svehla, (2012) **Vogel's Text book of Macro and Semimicro Qualitative Analysis**, Longman Inc., Newyork.	
2.	R. D. Madan, (2004) Modern **Inorganic Chemistry**, S. Chand & Co., New Delhi.	

3.	M.K. Jain. S.C. Sharma, (2004) **Modern Organic Chemistry**, Vishal publishing Co., New Delhi.
4.	B. R. Puri, L. R. Sharma, K. K. Kalia, (2014) **Principles of Inorganic Chemistry**, Milestone Publishers and Distributors, New Delhi.
5.	Arun Bahl, B.S.Bahl, (2012) **Advanced Organic Chemistry,** S. Chand & Co., New Delhi, Revised multicolor edition.
6.	Arun Bahl and B.S.Bahl, G.D.Tuli, (2012) **Essentials of Physical Chemistry**, S. Chand & Co., New Delhi, Revised multicolor edition.
REFERENCE BOOKS:	
1.	Sathya Praksash, G.D. Tuli, S. K. Basu, R.D. Madan, **(**2012) **Advanced Inorganic Chemistry,** Vol. 1, S. Chand & Co., New Delhi.
2.	J. D. Lee, (2006) **Concise Inorganic Chemistry**, Black Well Science, UK.
3.	M. K. Jain, S. C. Sharma, (2011) **Modern Organic Chemistry**, Vishal Publishing Co., New Delhi,.
4.	S. Glasstone, D. Lewis, (2004) **Elements of Physical Chemistry**, Macmillan Ltd, London.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	H
CO2	S	S	M	M	M
CO3	S	S	M	M	M
CO4	S	S	H	M	M
CO5	S	S	H	M	M
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – II** **INORGANIC, ORGANIC AND PHYSICAL CHEMISTRY -II**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	II	6	90	5

Course Objectives

1	Know about metallurgy, importance of periodic table and atomic properties.
2	To learn about Benzene and Aromaticity.
3	To study the fundamentals of thermodynamics and thermochemistry.

Course Outcomes (CO)

K1 – K5	CO1	Obtain problem solving skills in order to modify industrial processes in Extraction metallurgy.
	CO2	Gain knowledge about periodic properties
	CO3	Study of Aromatic Compounds and mechanism of certain reactions
	CO4	Learn about concepts of thermodynamics.
	CO5	Acquire the knowledge in thermochemistry.

UNIT – I	**Metallurgy**	**18 Hours**

***Introduction-Minerals, ores - occurrence of metals-classification of ores**-various steps involved in the metallurgical processes- concentration of ores, working of concentrated ore-Calcination – roasting - smelting - Thermodynamic Principles-reduction by carbon-Alumino Thermic process-reduction by heating air-electrolytic reduction - purification of metals- electro refining - zone refining - van Arkel process - Amalgamation process - Mond's process, Cement processing and mining operations.

UNIT-II	**Periodic table and atomic properties**	**18 Hours**

Mendeleev Periodic Table, Modern Periodic table-Long form periodic table, Description of groups-Alkali metals-Alkali earth metals-Boron Family-Carbon Family-Nitrogen family/ Pnicogens-Oxygen family/ Chalcogens-Halogen family-Inert gases/Noble gases –IUPAC nomenclature for the super heavy elements, Merits and Demerits of long form of periodic table.

Periodicity, Causes of periodicity, Calculation of effective nuclear charge, Periodic Properties, Atomic Size-Covalent radius--Ionic Radius-Metallic radius-Vander Waal's radius, Factors affecting atomic size, Ionization energy – factors affecting ionization energy-applications of ionization energy, Electron affinity-Factors affecting electron affinity, Electro negativity-Factors Affecting electro negativity-applications of electro negativity.

UNIT-III	**Aromatic Compounds**	**18 Hours**

Nomenclature of Aromatic Compounds – Mono substituted Benzene – Di substituted Benzene –

Polysubstituted Benzene - Concept of aromaticity –definition, Huckel's rule – application to Benzenoids and Non – Benzenoids (Cyclopentadienyl anion and tropylium cation). Preparations of Benzene: From acetylene, phenols and benzene carboxylic acids. General mechanism of electrophilic substitution, mechanism of nitration, sulphonation and halogenation, Friedel Craft's alkylation and acylation mechanism and their Limitations. Orientation of aromatic substitution - Definition of ortho, para, and meta directing groups. Ring activating and deactivating groups with examples. Orientation – (i) activating groups: Amino, methoxy and alkyl groups. (ii) Deactivating groups - carboxy, nitro, nitrile, carbonyl and sulphonic acid & halo groups. Synthesis of *p*- Bromo nitrobenzene and *o*-Nitrotoluene from benzene.

UNIT-IV	**Thermodynamics – I**	**18 Hours**

Basic concepts - scope and limitations - Thermodynamic terms - intensive and extensive properties-state, equilibrium - processes-nature of heat and work – pressure – volume work -isothermal reversible and irreversible expansion works of an ideal gas - maximum work - Zeroth law of thermodynamics - Internal Energy and First law of thermodynamics - Enthalpy of a system-Relation between ΔE and ΔH-Relation between Cp and Cv – Joule Thomson effect -comparison between adiabatic and isothermal expansions - Adiabatic expansion of an ideal gas -work done in adiabatic reversible expansion.

UNIT-V	**Thermo Chemistry**	**18 Hours**

Introduction - exothermic and endothermic reactions - Thermo chemical equations - Kirchoff's equation - types of heat of reaction - heat of formation - standard heat of formation - standard heat of reaction - heat of combustion-heat of solution - heat of neutralization-heat of fusion - heat of vaporization - heat of sublimation and heat of transition – definition - simple problems. Hess's law of constant heat summation – applications - bond energy and strength of bond - experimental determination of heat of combustion - bomb calorimeter

Self study portion

Teaching Methods: Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS:

1.	Puri B.R, Sharma L.R, Kalia K.C., (2014) **Principles of Inorganic Chemistry**, Milestone publishers and Distributors, New Delhi.
2.	Sathya praksash, G.D. Tuli, S. K. Basu, R.D. Madhan, (2012) **Advanced Inorganic Chemistry,** Volume 1, S. Chand & Company, New Delhi.
3.	M.K. Jain, S.C. Sharma, (2011) **Modern Organic Chemistry**, Vishal Publishing Co., Delhi.
4.	B.S. Bahl, and Arun Bahl, (2012) **Advanced Organic Chemistry**, S.Chand and Co, New Delhi, Revised multicolor edition.
5.	B.S. Bahl and G. D. Tuli, and Arun Bahl, (2012) **Essentials of Physical Chemistry,** S. Chand publishing, Revised multicolor edition.

REFERENCE BOOKS:	
1.	R. D. Madan, (2011) **Modern Inorganic Chemistry**, S.Chand and Co., Third Revised Edition.
2.	R. Gopalan, (2009) **Inorganic Chemistry For Undergraduate**, Universities Press (India) Private Limited, Hyderabad.
3.	Puri B. R. Sharma L. R., M. S. Pathania, **(**2013) **Principles of Physical Chemistry,** Vishal Publishing Co., New Delhi.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	H	H
CO2	S	S	M	M	M
CO3	S	S	H	H	M
CO4	S	S	H	H	M
CO5	S	S	H	H	M
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Programme Code: 04		**B.Sc.** Chemistry		
Title of the paper		**CORE PRACTICAL – I INORGANIC QUALITATIVE ANALYSIS AND PREPARATIONS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	II	3	90	2

Course Objectives

1.	To demonstrate the basic laboratory technique of semi micro qualitative analysis.
2.	To understand about the interfering anions, its elimination and group separation.
3.	To prepare inorganic complexes.

Course Outcomes (CO)

K1 – K5	CO1	Build the knowledge in principles of semi micro qualitative analysis.
	CO2	Know about the interfering and non-interfering anions.
	CO3	Experience to remove interfering anion and group separation of various cations.
	CO4	Group separation of various cations
	CO5	Learn the preparation of inorganic complexes.

I	**Semi - Micro Qualitative Analysis**
Analysis of a mixture containing two cations and two anions of which one will be an interfering ion. Semi-micro methods using the conventional scheme with Sodium Sulphide (Na2S) may be adopted.	

II	**Preparation of Inorganic Complexes.**
1	Tetraamminecopper(II) sulphate
2	Potassiumtrioxalatochromate(III)
3	Iron(III) hexacyanoferrate(II)
4	Hexamminecobalt(II) chloride
5	Hexathiourealead(II) nitrate
6	Tristhioureacopper(I) Sulphate
III	**Melting point** and **FT-IR spectrum** of Inorganic complexes.(demonstration only)

TEXT BOOKS:

1. Venkateswaran. V, Veeraswamy. R, Kulandaivelu . A.R, (1997), **Basic Principles of Practical Chemistry**, New Delhi, Sultan Chand and Sons,

REFERENCE BOOKS:

1.	G. Svehla, (1987), **Vogel's Qualitative Inorganic Analysis**, Orient Longman Ltd, Hyderabad.

Time: 3 Hours		**Max. Marks: 50**	
Distribution of Total Marks: 50			
	Record	5 marks	
	Analysis	40 marks	
	Preparation	5 marks	
Inorganic Analysis Marks: 40			
	Four ions with correct Procedure	40 marks	
	Three ions with correct Procedure	30 marks	
	Two ions with correct Procedure	25 marks	
	One ions with correct Procedure	12 marks	

- ❖ Spotting of an ion -5 marks
- ❖ Precipitation- 5 marks
- ❖ Correct detection of cations without eliminating the interfering ion should be treated as spotting
- ❖ At least one confirmatory test for each ion is expected. If no confirmatory test is reported deduct 3 marks
- ❖ The candidate may be asked to leave a small portion of the given mixture so that the examiners may confirm the presence of an ion if any discrepancies arise.
- ❖ Anions to be given: Sulphate, Chloride, Bromide, Fluoride, Borate, Nitrate, Carbonate, Oxalate, and Phosphate.
- ❖ Cations to be given: Lead, Bismuth, Copper, Manganese, Aluminium, Ferrous, Ferric, Cadmium, Cobalt, Nickel, Zinc, Barium, Calcium, Strontium, Magnesium and Ammonium.

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	S	H	M	M
CO2	S	S	H	M	M
CO3	S	S	H	M	M
CO4	S	S	S	M	M
CO5	S	S	S	H	H

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04		**B.Sc.** Chemistry		
Title of the paper		**CORE PAPER – III** **INORGANIC, ORGANIC AND PHYSICAL CHEMISTRY - III**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	III	4	60	4

Course Objectives

1.	To know the basic concepts in quantitative analysis.
2.	To observe the chemistry of dicarboxylic acids and reactions involving carbonyl compounds.
3.	To enumerate second law of thermodynamics, state functions S, A, G and chemical equilibrium.

Course Outcomes (CO)

K1 – K5	CO1	Gain knowledge in preparation, standardization of solution and principles of volumetric analysis.
	CO2	Study the preparation, properties and reactions of di carboxylic acids, unsaturated acids and hydroxy acids.
	CO3	To Study on the preparation and properties of aldehydes and ketones.
	CO4	Analyze and apply laws of thermodynamics.
	CO5	To understand the importance of absolute zero

UNIT – I	**Quantitative Analysis**	**12 Hours**

1. The mole concept – atomic, molecular and molar masses. Equivalent mass – Equivalent mass of an acid, equivalent mass of a base, equivalent mass of oxidizing and reducing agents.
2. Concentration terms – Normality, molarity, mole fraction, molality, percentage solution – weight composition, volume composition.
3. Principles of volumetric analysis – standard solution (primary and secondary standards) titration – types (acid, base, oxidation, reduction), equivalent point, end point, indicators – action of phenolphthalein and methyl orange, caution in volumetric titrimetry – precautions to avoid errors in titrimetric analysis, corrections for unavoidable errors.

UNIT-II	**Dicarboxylic acids**	**12 Hours**

Preparation, physical, chemical properties and uses of Oxalic, Malonic, Succinic, Glutaric, Adipic, Maleic and Fumaric acid. Geometrical isomerism of Maleic acid and Fumaric acid. Synthetic applications of Malonic ester and Acetoacetic ester, Tautomerism of Acetoacetic ester.

UNIT-III	**Aldehydes and ketones**	**12 Hours**

General methods of preparations and properties of aldehydes and ketones. Mechanism of

Nucleophilic addition of Grignard reagents, aldol condensation, Perkin, Knoevenagel, Claisen, Reformatsky reaction. Reactions with LiAlH4 and NaBH4, Wolf Kishner, Meerwein - Ponndorf-Verley reductions and Cannizzaro reactions, Clasein and Dickmann reactions.

UNIT-IV	**Thermodynamics-II**	**12 Hours**

Limitations of First law – need for second law. Various statements of second law. Spontaneous or irreversible processes. Criteria of spontaneity - Cyclic process- Entropy – definition - numerical definition. The Carnot's cycle- thermodynamic efficiency. Derivation of entropy from Carnot's cycle. Physical significance of entropy- Entropy change in isothermal expansion of an ideal gas, entropy change in reversible and irreversible processes, entropy change accompanying change of phase. Entropy as the function of P, V and T. Some other state functions-Work and Free energy functions - Gibb's free energy (G) and Helmholtz free energy (A)- variation of free energy with temperature and pressure. Isothermal change in free energy - Gibbs-Helmholtz equations. The Clapeyron Equation-Clausius – Clapeyron equation-its applications. Significance of ΔA and ΔG. van't Hoff isotherm –Van't Hoff isochore.

UNIT-V	**Thermodynamics-III**	**12 Hours**

Partial molal properties: Concept of chemical potential- Physical significance-The Gibbs –Duhem equation-variation of chemical potential with T and P- Time's Arrow .

Need for third law of thermodynamics – Nernst heat theorem- Third law of thermodynamics. Determination of absolute entropy of solids, liquids and gases. Standard absolute entropies (S^o). Entropy change in chemical reactions. Unattainability of absolute zero. Derivation of Boltzmann entropy equation. Residual entropy. Problems involving residual entropy. Stirlings approximation.

Self study portion

Teaching Methods: Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS:

1.	R. D. Madan, (2013) **Modern Inorganic Chemistry**, S. Chand & Co., New Delhi.
2.	B. R. Puri, L. R. Sharma, K. K. Kalia, (2017) **Principles of Inorganic Chemistry**, Milestone Publishers and Distributors, New Delhi.
3.	Arun Bahl, B. S. Bahl, (2010) A**dvanced Organic Chemistry**, S. Chand & Co., New Delhi.
4.	Arun Bahl and B. S. Bahl, G. D. Tuli, (2009) **Essentials of Physical Chemistry**, S. Chand & Co., New Delhi.

REFERENCE BOOKS:

1.	Sathya Praksash, G.D. Tuli, S. K. Basu, R.D. Madan, (2012) **Advanced Inorganic Chemistry**, Vol. 1, S. Chand & Co., New Delhi.

2.	J. D. Lee, (2006) **Concise Inorganic Chemistry**, Black Well Science, UK.
3.	M. K. Jain, S. C. Sharma, (2011) **Modern Organic Chemistry**, Vishal Publishing Co., New Delhi.
4.	B. Mehta, Manju Mehta, (2005) **Organic Chemistry**, Prentice Hall of India Pvt Ltd., New Delhi,
5.	S. Glasstone, D. Lewis, (2012) **Elements of Physical Chemistry**, Macmillan Ltd, London.
6.	B. R. Puri, L. R. Sharma, M. S. Pathania, (2009) **Principles of Physical Chemistry**, S. Chand & Co., New Delhi.

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	S	S	M	H
CO2	S	S	S	H	H
CO3	S	S	S	H	H
CO4	S	S	H	H	M
CO5	S	S	H	H	M

S – Strong **H** – High **M** – Medium **L** – Low

Subject code : 23UCH404

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – IV** **INORGANIC, ORGANIC AND PHYSICAL CHEMISTRY - IV**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	IV	4	60	4

Course Objectives

1.	To learn group IA elements.
2.	To know about various types of alcohols, phenols and their reactions
3.	To know about phase rule and phase equilibria

Course Outcomes (CO)

K1 – K5	CO1	Gain the knowledge about the properties of alkali metals.
	CO2	Understand the basic aspects of phenols, amines and its derivatives.
	CO3	Analyze and apply phase rule to various systems.
	CO4	Understand colligative properties and their determinations.
	CO5	Understanding Intellectual Properties and the importance of it and awareness of patents

UNIT – I	**Group IA elements - Alkali metals**	**12 Hours**

***General properties-electronic configuration, density, molar volume, atomic volume, atomic and ionic radii, heat of atomization**, melting and boiling point- ionization energy-Comparison of ionization energy of alkali metals- electropositive character. Similarities in chemical properties – formation of halides, nitrates, oxides, hydroxides, hydrides. Comparison of Lithium with other members of the family. Anomalous behavior of lithium- diagonal relationship of Li and Mg - Lithium- Extraction, properties and uses. Sodium extraction, properties, commercial and analytical uses

UNIT-II	**Alcohols**	**12 Hours**

Classification and Nomenclature, **Monohydric Alcohols :** Methods of Formation by Reduction of Aldehydes, Ketones, Carboxylic Acids and Esters, Hydrogen Bonding, Acidic Nature, Reactions of Alcohols. **Dihydric Alcohols:** Methods of Formation, Chemical Reactions of Vicinal Glycols, Oxidative Cleavage [$Pb(OAc)_4$ and HIO_4] **Trihydric Alcohols :** Methods of Formation, Chemicals Reactions of Glycerol. Commercial importance of methanol in energy field

UNIT-III	**Phenols**	**12 Hours**

Monohydric Phenols - preparation and properties, acidity of phenols, reaction of monohydric phenols - Esterification, Nitration, Sulphonation, Halogenation, coupling with diazonium salts, Kolbe-Schimit, Reimer-Tiemann, Schotten-Baumann, Hoesch-Houben reaction and Gattermann reactions with mechanism, Lederer Manasse reaction, Phenol formaldehyde resin, Phenolphthalein preparation.

UNIT-IV	**Phase rule and Phase equilibria**	**12 Hours**

Phase rule-statement- Definition of terms phase, component, and degrees of freedom. Conditions for equilibrium between phases, Derivation of Gibb's Phase rule. Phase diagrams of Water, Carbon dioxide and Sulphur systems. Polymorphism- transition temperature-cooling curves-Two components system: Simple Eutectic systems - Silver–Lead and Formation of compounds with congruent and incongruent melting point (Ferric chloride – water system and Sodium sulfate - water system).

UNIT-V	**Intellectual Property Rights**	**12 Hours**

Introduction: Introduction to Intellectual property, types of intellectual property, importance of intellectual property rights, agencies Responsible for Intellectual property Registration, Regulatory – Compliance and Liability Issues.

Patents: Law of patent - Advantages of Patents - Patent Searching Process - The Need for a Search - Searching Methods. Patent Application Process - Overview of the Application Process. Patent Practice - Confidentiality of Application Process and Publication of Patent Application. Types of Application - Preparing the Application. Patent Prosecution Flowchart - Ownership Rights - Ownership transfer.

Self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	R. D. Madhan, (2012)**Modern Inorganic Chemistry**, S. Chand and Co.
2.	Arun Bahl and B.S.Bahl, (2010) **Advanced Organic Chemistry**, S. Chand and Co., New Delhi.
3.	B. R. Puri, L.R. Sharma, and S. Pathania, (2019) Principles of Physical Chemistry, Shoban Lal Nagin Chand & Co, New Delhi.
4.	Sathya Praksash, G.D. Tuli, S. K. Basu, R.D. Madhan, (2012) Advanced Inorganic Chemistry, Volume 1, S. Chand & Company, New Delhi.
5.	**Intellectual property right**, Deborah, E. Bouchoux, cengage learning publications.
6.	**Intellectual Property Rights**: N K Acharya:, Asia Law House; 6th edition

REFERENCE BOOKS:

1.	Lee J.D., (2007) **Concise Inorganic Chemistry**, Black Well Science-Wiley – India, New Delhi.
2.	B.Mehta, Manju Mehta, (2005) **Organic Chemistry**, Prentice Hall of India Private Limited, New Delhi.
3.	Arun Bahl, B.S. Bahl, G.D.Tuli., (2009) **Essentials of Physical Chemistry** S. Chand and company, New Delhi.

4.	C B Raju, **Intellectual Property Rights**: Serials Publications (2006).
5.	A. Subbian, **Intellectual Property Rights** – Heritage, Science, & Society under international treaties –Deep & Deep Publications – New Delhi.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	M	M	M
CO2	S	S	M	M	M
CO3	S	S	M	M	M
CO4	S	S	H	H	M
CO5	S	S	H	H	M
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PRACTICAL – II INORGANIC VOLUMETRIC AND ORGANIC QUALITATIVE ANALYSIS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**IV**	3	90	3

Course Objectives

1.	To demonstrate the concept of quantitative volumetric analysis.
2.	To understand the various types of titrimetric analysis.
3.	To identify the functional groups of unknown organic compounds.

Course Outcomes (CO)

K1 – K5	CO1	Gain the knowledge in principles of volumetric analysis.
	CO2	Estimating the amount of substances present in solutions.
	CO3	Learn to approach a problem systematically
	CO4	Interpret the results logically.
	CO5	Detect various functional groups present in an organic compound.

I	**Titrimetric Quantitative Analysis**

a. Acidimetry and Alkalimetry:

1.Estimation of HCl by NaOH using a standard Oxalic acid solution

2. Estimation of Na_2CO_3 by HCl using a standard Na_2CO_3 Solution.

b. Permanganametry:

1.Estimation of Oxlaic acid by $KMnO_4$ using a standard Oxalic acid solution

2.Estimation Iron(II) Sulphate by $KMnO_4$ using a standard Mohr's Salt solution

3.Estimation of Calcium(II) by $KMnO_4$ using standard oxalic acid solution

c. Dichrometry:

1.Estimation of Iron (II) by potassium dichromate using standard Mohr's salt solution

d. Iodometry:

1. Estimation of $KMnO_4$ by Thio using a standard Potassium dichromate Solution

2. Estimation of Copper (II) Sulphate by $K_2Cr_2O_7$ solution.

II	**Organic analysis**
	Analysis of organic compounds
1	Preliminary tests
2	Detection of elements present
3	Aromatic or Aliphatic
4	Saturated or Unsaturated

5	Nature of the functional group
6	Confirmatory tests and Preparation of derivatives for the functional groups.
7	Melting point and FT-IR spectrum of organic compounds **(demonstration only)**

The following functional group compounds may be given:
Aldehydes, Ketones, Amines, Amides, Diamide, Carbohydrates, Phenols, Acids, Esters and Nitro compounds.

TEXT BOOKS :

1.	Venkateswaran. V, Veeraswamy. R, Kulandaivelu. A.R, (1997), **Basic Principles of Practical Chemistry**, New Delhi, Sultan Chand and Sons,
2.	Mendham. J, Denney. R.C, Bames. J.D, and Thomas, M. (2006) Vogel's Text book of Quantitative Chemical Analysis, Pearson Education.

REFERENCE BOOKS:

1.	Gopalan. R, Subramaniam. P.S, and Rengarajan. K, (2010) Elements of Analytical Chemistry, Sultan Chand and Sons..

Time: 6 Hours		**Max. Marks : 50**	
Distribution of Total Marks: 50			
	Record	05 marks	
	Analysis	45 marks	
Distribution of Practical Marks-40			
	Volumetric	20 mark	
	Organic Analysis	25 mark	
Volumetric Analysis Marks-20			
	Procedure (To be written within five minutes)	5 mark	
	Experiment	15 mark	
	Error up to <2 %	15 mark	
	2 – 3 %	10 mark	
	3 - 4 %	5 mark	
	>4 %	3 mark	
Organic Analysis Marks- 25			
	Preliminary Tests	6 mark	
	Aliphatic or Aromatic	4 mark	
	Saturated or Unsaturated	3 mark	
	Special elements	4 mark	
	Functional group	5 mark	

	Derivative	3 mark	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	S	H
CO2	S	S	H	H	M
CO3	S	S	H	H	H
CO4	S	S	H	S	H
CO5	S	S	H	S	H
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Subject code : 23UCH505

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – V SPECTROSCOPY AND CHROMATOGRAPHIC TECHNIQUES**		
Batch	**Semester**	Hours / Week	Total Hours	Credits
2023 - 2024	V	3	45	3

Course Objectives

1.	To know about the region of electromagnetic spectrum, fundamentals of ultra – violet visible spectroscopy and Infrared spectroscopy.
2.	To study Nuclear Magnetic Resonance (NMR) spectroscopy and Mass
3.	Spectrometry and to interpret and solve problems using various spectra.

Course Outcomes (CO)

K1 – K5	CO1	Understand the basic principles of UV-Visible spectroscopy and to utilize their basic aspects to identify various organic compounds.
	CO2	Gain the knowledge in principles, and functions of IR spectroscopy.
	CO3	Study the basic principles of NMR spectroscopy and apply to identify the organic compounds.
	CO4	Know about basic principles of mass spectroscopy technique and the application of various spectral techniques to elucidate the structure of organic molecules.
	CO5	Exploring the various chromatography techniques and their applications in separation of organic mixtures.

UNIT – I	**Ultra – violet and visible spectroscopy**	**9 Hours**

Electromagnetic spectrum and absorption of radiations, Principle of ultraviolet spectroscopy, the absorption laws- Lambert's law and Beer's law. Selection rules, instrumentation – Block Diagram, theory of electronic spectroscopy, types of electronic transitions, the chromophoric concept, auxochromes, absorption and intensity shifts – bathochromic, hyPSO chromic hyperchromic and hypochromic shifts. Types of absorption bands, solvent effects, Frank – Condon principle. Woodward fisher rules for calculating absorption maximum in dienes.

UNIT-II	**Infrared spectroscopy**	**9 Hours**

Principle of Infrared spectroscopy, molecular vibrations, vibrational frequency, number of fundamental vibrations, overtones and combination bands, selection rules, factors influencing vibrational frequency – coupled vibrations and Fermi resonance, electronic effects, hydrogen bonding and bond angles. Scanning of infrared spectrum (instrumentation), finger print region. only.

UNIT-III	**Nuclear Magnetic Resonance (NMR) spectroscopy**	**9 Hours**
Principle and theory of NMR spectra, conditions of resonance, relaxation process – spin –spin relaxation, spin – lattice relaxation and quadrupole relaxation. Instrumentation, solvent used in NMR, Chemical shift (shielding and deshielding effects). Number of signals, position of signals, factors influencing chemical shift – inductive effect, vander-Waal's deshielding, Anisotropy (diamagnetic shielding in benzene), Peak area and proton coupling, spin – spin splitting, splitting of the signals in pure ethanol and chloroethane.		
UNIT-IV	**Mass spectrometry and application of spectroscopic methods**	**9 Hours**
Basic principles, theory of mass spectrometry, meta stable ions or peaks, nitrogen rule, general fragmentation modes of hydrocarbons, Retro – Diels Alder reaction, hydrogen transfer rearrangements and Mclafferty rearrangement. ***IR, NMR and Mass techniques in the identification of simple organic molecules.** (Ethanol and dimethyl ether, acetaldehyde and acetone, ethylene and acetylene).		
UNIT-V	**Chromatography**	**9 Hours**
Introduction, definition, types, thin layer chromatography – experimental techniques - coating materials - preparation of thin layer in plates – activation of adsorbents – purification of silica gel – sample application – development tank – solvent systems – plate development – detection of components, evaluation of chromatography. Column chromatography – principle experimental techniques – apparatus – adsorbents – preparation of adsorption columns – solvents used in successive elution – gradient elution.		
Self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Y.R, Sharma, (2007) **Elementary Organic Spectroscopy Principles and Chemical Applications**, S.Chand & Company Ltd, New Delhi.	
2.	Gurdeep R. Chatwal, (2018) **Instrumental Methods of Chemical Analysis**, Himalaya publishing house, Delhi.	
3.	V.K. Srivastava, K.K. Srivastava, K.K. Kishore, **Introduction to Chromatography – Theory** & Practice, S. Chand & Co. (P) Ltd., New Delhi, 3rd Edn., (2010).	
REFERENCE BOOKS:		
1.	P.S. Sindhu, **Elements of (2010) Molecular Spectroscopy**, New Age International	

	Publishers, New Delhi.
2.	H.S. Randhana, (2007) **Modern Molecular Spectroscopy**, Macmillan India Ltd, New Delhi.
3.	H. Kaur, **Instrumental Methods of Chemical Analysis**, Pragati Prakashan, Meerut, 2010.
4.	Jag Mohan, (2018) **Organic Spectroscopy – Principles and Applications**, Narosa publishing house
5.	A.K. Srivastava, P.C. Jain, **Chemical Analysis: An Instrumental Approach** for B.Sc. Hons. and M.Sc. Classes, S. Chand and Company Ltd., Ram Nagar, New Delhi, 2010.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	H	M
CO2	S	S	M	M	M
CO3	S	S	H	M	H
CO4	S	S	S	H	H
CO5	S	S	S	H	H
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Subject code : 23UCH506

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – VI INORGANIC CHEMISTRY**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	V	3	45	3

Course Objectives

1.	To understand the key features of coordination compounds, including: the variety of structures, ligands, various theories of coordination complexes, stability of complexes.
2.	To identify what radioisotopes and acquaint knowledge about nuclear reactions.
3.	To describe about Inorganic acids, bases, Inorganic Solvents and Inorganic Polymers.

Course Outcomes (CO)

K1 – K5	CO1	Understand the theories of co-ordination compounds.
	CO2	Knowledge about basics nuclear Chemistry
	CO3	Analyze the importance of radioactive isotopes and nuclear reactions.
	CO4	Describe about the different concepts of Inorganic acids, bases, Inorganic Solvents and Inorganic Polymers.
	CO5	Understanding the importance of Inorganic Solvents and Inorganic Polymers.

UNIT – I	**Co-ordination Compounds**	**9 Hours**

Co-ordination Compounds – Types of ligand, Industrial applications of Chelation, Nomenclature, Werner's coordination theory- Electronic interpretation of coordinate bond by Sidgwick. EAN rule, Isomerism: Examples of geometrical isomerism in square planar and octahedral coordination compounds. Magnetic properties of square planar and octahedral coordination compounds and their interpretation by Pauling's Valence Bond Theory.

UNIT-II	**Nuclear Chemistry - I**	**9 Hours**

Introduction – Nuclear stability and n/p ratio. Magic numbers. Packing fraction. Mass defect and binding energies. Definition for isotopes, isobars and isotones. Detection of isotopes - mass spectrographs-Thomson's, Dempster's and Aston's mass spectrographs - Importance of discovery of isotopes. Radioactivity – emission of alpha, beta and gamma rays. Radioactive disintegration - first order kinetics - half-life period. Radioactive disintegration series. Basic difference between nuclear reaction and conventional chemical reaction.

UNIT-III	**Nuclear Chemistry - II**	**9 Hours**

Nuclear reactions – fission, fusion, spallation, capture and particle-particle reactions - nuclear fission-nuclear reactors- basic components of nuclear reactor, Atom bomb- nuclear fusion - Stellar energy- Hydrogen bomb. Artificial transmutation of elements. Artificial radioactivity. Uses of radioactive isotopes – medicine – agriculture – C^{14} dating – dating of Universe. ***Atomic power**

projects in India- Disposal of nuclear wastes, Cyclotron.		
UNIT-IV	**Acids and Bases**	**9 Hours**
Different concepts of acids and bases- Arrhenius, Lowry- Bronsted, Lewis, Cady- Esley (solvent system), Lux-Flood and Usanovichs concepts. Conjugate acids and bases- comparison of strengths of Lowry-Bronsted acids - Comparison of strengths of Lewis acids and bases. Levelling effect and levelling solvents. Hard and soft acids and bases (HSABs)- Acid and bases strengths of HSABs- Applications of HSABs concept, Basis of hardness and softness - pi-bonding contributions- electro negativity factor. Limitations of HSAB concept.		
UNIT-V	**Solvents and Inorganic Polymers**	**9 Hours**
Solubilities of compounds - effect of temperature on solubility- chemical structure and solubility. Classification of solvents - properties of ionizing solvents. Types of reactions in solvents. Specific non-aqueous solvents - protic solvents (ammonia) - aprotic solvents (SO_2). Types of inorganic polymers, Classification, properties of inorganic polymers, synthesis, structural aspects and applications of silicones, types(high thermal, resins, fluids, rubbers, greases) and synthesis, structural aspects and applications of Borazines and substituted borazine -Boroxine, N-trimethyl borazine- B- trimethyl borazine.		
Self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	R. D. Madhan, G. D. Tuli, and S. M. Malik,(2009) Selected Topics in Inorganic Chemistry, S. Chand & Co., New Delhi.	
2.	B. R. Puri, L. R. Sharma, K. K. Kalia, (2019) **Principles of Inorganic Chemistry**, Milestone Publishers and Distributors, New Delhi	
3.	U. N. Dash, (2010) **Nuclear Chemistry**, S. Chand & Co., New Delhi U. N. Dash, (2010) Nuclear Chemistry, S. Chand & Co., New Delhi	
REFERENCE BOOKS:		
1.	Karen C.Timberlake,(2008) **Basic Chemistry**, Los Angeles Valley College, Pearson-Benjamin Cummings, San Francisco, First Edition.	
2.	G. S. Manku,(2006), **Theoretical Principles of Inorganic Chemistry**, Mc - Graw Hill Edition.	
3.	M. G. Arora, M. Singh, (2014) **Nuclear Chemistry**, Anmol publications Pvt. Ltd., New Delhi.	
4.	Sathya Praksash, G.D. Tuli, S. K. Basu, R.D. Madhan,(2012) **Advanced Inorganic Chemistry**, Vol. 1, S. Chand & Co., New Delhi.	
5.	H. J. Arnikar,(2018) **Essentials of Nuclear Chemistry**, New Age International.	

6.	R. D. Madan, (2019) **Modern Inorganic Chemistry**, S. Chand & Co., New Delhi.

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	S	H	M	H
CO2	S	M	S	M	S
CO3	S	M	S	M	S
CO4	S	S	S	S	S
CO5	S	S	S	S	S

S – Strong **H** – High **M** – Medium **L** – Low

Subject code : 23UCH507

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – VII** **ORGANIC REACTION MECHANISM**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**V**	4	60	3

Course Objectives

1.	To study asymmetry and optical activity of organic molecules and basics in carbohydrate.
2.	To understand the mechanisms of important organic rearrangements reactions and Preparations and reactions of Amines and Diazo compounds.
3.	To study preparation and properties of heterocyclic compounds.

Course Outcomes (CO)

K1 – K5	CO1	Understanding the fundamental aspects of stereochemistry.
	CO2	Learn about preparation, properties and structural elucidation of carbohydrates.
	CO3	Study on the various naming reactions and their detailed mechanistic pathway.
	CO4	Acquire the knowledge about the preparations and reactions of Amines and Diazo compounds.
	CO5	To inculcate knowledge about five and six membered heterocyclic compounds

UNIT – I	**Stereochemistry**	**12 Hours**

Optical Isomerism, cause of optical activity, plane polarized light, specific rotation, Plane of symmetry, chiral (asymmetric) carbon atom, chirality, Optical isomerism of lactic acid, Fischer projections and optical isomerism of tartaric acid. Properties of Enantiomers and diastereo isomers. Resolution of Racemic mixture – mechanical separation – kinetic separation – selective adsorption – chemical method – biochemical method. Racemization, chiral (Asymmetric) synthesis, Walden inversion. Specifying absolute configuration – D, L and R, S system for asymmetric molecule. Optical activity of Biphenyl, Allenes, Spiranes and overcrowded molecules. Geometrical isomerism.

UNIT-II	**12 Hours**

Carbohydrates - *Introduction, classification

Monosaccharides - occurrence, preparation, structural elucidation, properties and uses of Glucose and Fructose. Cyclic form of glucose and fructose. Mutarotation, interconversion of glucose to fructose and vice versa.

Disaccharides – Structure, preparation, properties and uses of sucrose and maltose.

Polysaccharides – Starch and cellulose - Manufacture, structure and properties. Derivatives of cellulose.

UNIT-III	**Molecular Rearrangements with mechanism**	**12 Hours**
Reaction, mechanism, and applications of molecular rearrangement reactions - Pinacol-Pinacolone, Beckmann, Hoffmann, Curtius, Benzilic acid, Claisen Rearrangements, Cope rearrangement and Fries rearrangement.		
UNIT-IV	**Amines and Diazo compounds**	**12 Hours**
Preparation and reactions of amines, separation of a mixture of primary, secondary and tertiary amines – comparison of their basicity. Hinsberg test for distinguishing primary secondary and tertiary amines, Ring substitution, Diazotization and coupling reaction of aromatic amines. Preparation, structure and synthetic applications of Diazomethane.		
UNIT-V	**Heterocyclic compounds**	**12 Hours**
Introduction- preparation – physical and chemical properties of Furan, Pyrrole, Thiophene, Pyridine, Quinoline and Isoquinoline.		
Self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		

TEXT BOOKS :

1.	Arun Bahl and B.S.Bahl, (2012) **Advanced Organic Chemistry**, S. Chand and Co., New Delhi.
2.	Gurdeep. R. Chatwal, (2017) **Reaction Mechanism and Reagents in Organic Chemistry**, Himalaya Publishing House Delhi.
3.	M.K. Jain, S.C. Sharma, (2013) **Modern Organic Chemistry**, Vishal Publishing Co., New Delhi.

REFERENCE BOOKS:

1.	M.G Arora, (2008) **Stereochemistry in Organic Compounds**, Anmol Publications Private Ltd New Delhi.
2.	Jagdamba Singh and Yadav, (2009) **Organic Synthesis**, Vol. I and II. Pragathi and Prakasam Publishers.
3.	I.L.Finar,(2009) **Organic Chemistry,Vol.I and II**, Addison-Wesley Longman.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	M	S	M	M
CO2	S	S	S	S	S
CO3	S	S	S	M	M
CO4	S	M	S	M	M
CO5	S	H	S	H	H
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER - VIII PHYSICAL CHEMISTRY - I**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	V	4	60	4

Course Objectives

1.	To understand the fundamentals of electrochemistry.
2.	To know the types and importance of electrodes and electro chemical cells.
3.	To study about corrosion, batteries and Electroanalysis.

Course Outcomes (CO)

K1 – K5	CO1	Understanding the concept of conductance and its applications.
	CO2	Acquire basic knowledge about electrode potential, electrochemical cell and potentiometric titrations.
	CO3	Understanding the fundamental principles of corrosion, protective coatings
	CO4	electroplating and its significance.
	CO5	Know about basic principles and instrumentation of Electrochemical Power Systems, Polarography and its applications.

UNIT – I	**Fundamentals of Electrochemistry**	**12 Hours**

Introduction-Classification of conductors – Electrolytic conductance-conductivity cell measurement of conductance of solutions – Variation of equivalent conductance with dilutions. Migrations of ions - Transport number – determination by moving boundary method and Hittorf's method – Kohlrausch's law – statement - application. Arrhenius theory of electrolytic dissociation - Ostwald's dilution law and limitations - theory of strong electrolytes: Debye-Huckel – Onsagar theory (elementary treatment only) - Debye – Falkenhagen effect and Wien effect. Applications of conductance measurements: (i) Determination of dissociation constant of a weak organic acid (i) Conductometric titrations – acid-base titration, precipitation titration. (iii) Determination of solubility product of sparingly soluble salt (iv).Determination of ionic product of water.

UNIT-II	**Electro Chemical Cells**	**12 Hours**

Galvanic cell- classification of cell – representation of cell – cell terminology. Reversible and irreversible cells - Electrode potentials – types of electrodes and their potentials. Nernst equations - Computation and measurement of cell emf and Weston - Cadmium cell - Single electrode potentials. Determination and significance of electrode potentials - Electrochemical series and its applications- Thermodynamic quantities of cell reactions (ΔG, ΔH and ΔS).

UNIT-III	**Electrodes and their types**	**12 Hours**

pH scale – Buffer solution, Buffer action – Henderson's equation – determination of pH of Buffer

solution. Concentration cells with and without transport - Liquid junction potential. Application of EMF measurements – determination of pH using hydrogen, glass and quinhydrone electrode. Potentiometric titrations – acid - base, redox and precipitation. Redox Potentials - redox indicators - diphenyl amine.

UNIT-IV	**Corrosion and Adsorption**	**12 Hours**

*__Introduction, Dry or Chemical corrosion, Wet or Electrochemical Corrosion, Mechanism of Wet of Electrochemical Corrosion__, Galvanic (or Bimetallic) Corrosion, Concentration Cell Corrosion, Passivity, Types of Corrosion, Galvanic Series, Factors Influencing Corrosion, Corrosion Control Methods. Cathodic Protection Protective Coatings, Metallic Coating, Electroplating, Influencing factors, Pretreatment, Electroplating Methods – Cu, Ni, Electroless Plating – Cu, Ni.

Sorption – Absorption, Adsorption-Types of adsorption, adsorption of gases by solids. Adsorption isotherms –Freundlich, Langmuir. Adsorption of solutes from solutions. Application of adsorption.

UNIT-V	**Electrochemical Power Systems and Analysis**	**12 Hours**

*__Introduction – Batteries – Types__- Lead storage cells and Lithium ion cell. Fuel cells - Definition and importance, Hydrogen-Oxygen fuel cell, Hydrocarbon - Oxygen cell. Zinc air battery - Over voltage – Application of over voltage. Polarography – Instrumentation - Advantages of DME-Limiting current, factors affecting limiting current - Ilkovic equation (derivation not necessary) - Half wave potential –Application of polarography. Amperometric Titrations. Bio electrochemistry – Bio electrochemical cell - Electrochemical mechanism of Nervous System.

Self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	P. L. Soni. O. P. Dharmarha and U. N. Dash, (2013) **Textbook of physical chemistry**, S. Chand & Co., New Delhi.
2.	B.S. Bahl and G. D. Tuli, and Arun Bahl, (2012) **Essentials of Physical Chemistry**, S. Chand publishing, Revised multicolor edition.
3.	P. C. Jain and Monika Jain, (2013) **Engineering Chemistry**, Dhanpat Rai Publishing Co., New Delhi.

REFERENCE BOOKS:

1.	B. R. Puri, L. R. Sharma, M. S. Pathania, (2013) **Principles of Physical Chemistry**, S. Chand & Co., New Delhi
2.	Karen C. Timberlake, (2005) **Basic Chemistry**, Los Angeles Valley College, Pearson

	Benjamin Cummings New York.
3.	Samuel Glasstone,(2002) **Introduction to Electrochemistry**, EWP Pvt. Ltd.
4.	Syed Aftab Iqbal, (2011), **Text Book of Electrochemistry**, Discovery Publishing house Pvt. Ltd., New Delhi.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	S	S
CO2	S	S	H	M	H
CO3	S	S	H	M	H
CO4	S	S	M	M	H
CO5	S	H	M	S	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – IX** **SOLID STATE AND COORDINATION CHEMISTRY**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	VI	5	75	4

Course Objectives

1.	To know about fundamentals of crystallography and solid state Chemistry
2.	To study about reactions of complexes.
3.	To in sight knowledge about Bio – Inorganic Chemistry

Course Outcomes (CO)

K1 – K5	CO1	Knowing the difference between amorphous and crystalline solids and their arrangement in crystal lattice.
	CO2	Learn about defects in crystals, various theories of metallic bonding and alloys.
	CO3	Decide the various crystal structures using X-ray diffraction techniques and Study about liquid crystals.
	CO4	Study about various ligand substitution reactions.
	CO5	To acquire knowledge about bioinorganic chemistry.

UNIT – I	**Solid state – I**	**15 Hours**

Classification of solids (true solids, pseudo solids, crystalline and amorphous solids), elements of crystal symmetry, Definition of lattice point, crystal lattice and unit cell. Seven crystal system and 14 Bravais lattices, close packing of identical solid spheres (CCP, FCC, HCP and BCC), Designation of planes in crystals – Miller indices, radius ratio rule and shape of ionic crystal. Number of particles per unit cell and density of crystals.

UNIT-II	**Crystallography**	**15 Hours**

Study of the following with respect to cubic system: (100), (110) and (111) planes. Inter planar distances –d100, d110, d111- Ratio of inter planar distances ($d_{100} : d_{110} : d_{111}$) X-ray diffraction studies of crystals –Bragg's equation –Bragg method and powder method – crystal structure of NaCl and ZnS. Growth of a crystals from the Melt and the solution.

UNIT-III	**Solid state – II**	**15 Hours**

1. ***Defects in crystal** – Stoichiometric and non-Stoichiometric defects.
2. Metallic bonding- theories- electron gas theory, Valence bond theory, Molecular orbital theory (Band theory) –True metal or conductor, insulators, semiconductors-types of semiconductors-intrinsic and extrinsic, n and p- type.
3. Alloys-Substitutional and interstitial solid solutions, intermetallic compounds- Tamman's rule,

Hume-Rothery rule. Alloys in automobile industry and construction sectors.		
UNIT-IV	**Reactions mechanism of coordination compounds**	**15 Hours**
Ligand substitution reactions in octahedral complexes, labile and inert complexes – types and mechanism of substitution reactions – SN_1 and SN_2 type mechanisms – acid hydrolysis reaction – simple acid hydrolysis type and atalyzed aquation type, base hydrolysis reaction – SN_2 and SN_1CB mechanism – anation reactions. Ligand substitution reactions in square planar complexes – trans effect – trans directing series – theories of trans effect – applications of trans effect.		
UNIT-V	**Bioinorganic chemistry**	**15 Hours**
Porphyrin systems- Structure of myoglobin and hemoglobin. Role of Hemoglobin in biological systems – cooperativity effect – explanation of cooperativity effect in hemoglobin-metallo enzymes –Carbonic anhydrase, inhibition and poisoning of enzymes – role of alkali and alkaline earth metals in biological systems –sodium pump –calcium pump – biological functions and toxicity of some elements – biological fixation of nitrogen.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Sathya Prakash, G.D. Tuli, S. K. Basu, R.D. Madhan, (2015) Advanced Inorganic Chemistry, Volume 1, S. Chand & Company, New Delhi.	
2.	Wahid Malik, G.D Tuli, R. D. Madhan, (2015) Selected Topics in Inorganic Chemistry, S. Chand & Company, New Delhi.	
3.	B.R.Puri,L.R.Sharma and K.C.Kalia, (2009) Principles of Inorganic Chemistry Milestone Publishers, New Delhi.	
REFERENCE BOOKS:		
1.	M.G. Arora, (1997) **Solid State Chemistry**, Anmol Publishing House, New Delhi.	
2.	R.D.Madhan, (2011) **Modern Inorganic Chemistry**, S.Chand & Company, New Delhi	
3.	Gurdeep raj, (2011) Advanced **Inorganic Chemistry**, Vol.1 , Goel Publishing House, Meerut.	
4.	Gurdeep raj, (2010) **Advanced Physical Chemistry**, Vol.1 , Goel Publishing House, Meerut.	
5.	Asim K.Dass, (2007) **Bioinorangic Chemistry**, Books and Allied (p) Ltd, Kolkata.	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	M	H
CO2	S	S	H	H	H
CO3	S	S	S	M	M
CO4	S	M	S	S	S
CO5	S	H	S	M	H
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Subject code : 23UCH610

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – X** **CHEMISTRY OF NATURAL PRODUCTS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**VI**	5	75	4

Course Objectives

1.	To study about Terpenoids and Alkaloids.
2.	To understand about Vitamins and Hormones.
3.	To study the preparations and reactions of amines, Diazocompounds and Chemotheraphy.

Course Outcomes (CO)

K1 – K5	CO1	Study on the classification, structural elucidation and synthesis of few important terpenoids.
	CO2	Learn about structural determination and synthesis of alkaloids.
	CO3	Acquire basic knowledge about vitamins and hormones.
	CO4	To study about Amino acids, peptides and Proteins.
	CO5	To gain knowledge about chemotherapy.

UNIT – I	**Terpenoids**	**15 Hours**

Terpenoids: Introduction, classification, isolation and commercial value of essential oils. Isoprene rule, structural elucidation and synthesis of Geraniol, terpineol, dipentene and alpha-pinene.

UNIT-II	**Alkaloids**	**15 Hours**

Introduction, classification, phytochemical isolation of naturally occurring products, general characteristics and general methods of determining structures and Hoffmann's exhaustive methylation. Structural elucidation and synthesis of Nicotine, Coniine, Piperine and Papaverine.

UNIT-III	Vitamins and Hormones	**15 Hours**

***Introduction, classification, sources of Vitamins and their deficiency diseases**. Structural elucidation and synthesis of Thiamine and Riboflavin.

Hormones -Introduction, structural elucidation and synthesis of adrenaline and thyroxin.

UNIT-IV	**Amino acids, peptides and Proteins.**	**15 Hours**

1. **Amino acids** – Nomenclature, dipolar nature of amino acids, isoelectric point, methods of preparation – amination of halo acids – Strecker synthesis – Gabrielphthalimide synthesis – Koop synthesis. Physical and chemical properties of amino acids.

2. **Nomenclature of peptides** – peptide linkages, determination of structure of peptides – synthesis

of peptides.

3. **Proteins** - classification based on composition and function. Structure of proteins – primary – secondary and tertiary structure. General properties of proteins. Denature of proteins. Colour test for proteins.

UNIT-V	**Chemotherapy**	**15 Hours**

Designation of drugs based on physiological action – functional and formaco dynamic drugs – chemotherapeutic drugs. Definition and two examples each of antibacterial drugs- sulpha drugs and mode of action of sulpha drugs, Antimalarial drugs, Amebicidal drugs, Antiseptics, Anaesthetics, Analgesics – Narcotic and synthetic- Antipyretics and anti-inflammatory agents - Antibiotics - Penicillin, streptomycin, and tetracyclins.

self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	Arun Bahl and B.S.Bahl, (2015) Advanced Organic Chemistry, S. Chand and Co., New Delhi.
2.	M. K. Jain and S. C. Sharma, (2011) Modern Organic Chemistry, New Delhi

REFERENCE BOOKS:

1.	Jagdamba Singh and Yadav, (2015) Organic Chemistry, Vol. I and II, Pragathi Prakasam Publishers.
2.	I. L. Finar, (2010) Organic Chemistry, Vol. I and Vol.II, Addison-Wesley Longman.
3.	Gurdeep Chatwal, (2013) Organic Chemistry of Natural Products, Himalaya Publishing House, New Delhi.
4.	Morrison R.T. and Boyd R.N, (2008) Organic Chemistry New York, Allyn & Bacon Ltd.

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	H	S	M	H
CO2	S	H	H	M	H
CO3	S	H	M	H	S
CO4	S	S	S	S	S
CO5	S	S	H	S	H

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PAPER – XI PHYSICAL CHEMISTRY - II**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**VI**	4	60	4

Course Objectives

1.	To understand the basics and theoretical aspects of Chemical kinetics.
2.	To learn about kinetics of thermal and photochemical reactions.
3.	To gain knowledge about importance of catalysis, colloids and Liquid state.

Course Outcomes (CO)

K1 – K5	CO1	Understand the basic principles, various experimental techniques and Theories of chemical kinetics.
	CO2	To understand the importance of various theories explaining chemical kinetic.
	CO3	Gain the knowledge about principles of photochemical and photosensitized Process.
	CO4	Study the basic principles and types of catalysis and colloids.
	CO5	Explore the fundamentals of Liquid State.

UNIT – I	**Chemical Kinetics-I**	**12 Hours**

Empirical laws and experimental aspects. Rate law, stoichiometry, order and molecularity of reactions. Setting up and solving simple differential equations for first order, second order, third order and zero order reaction. Expressions for half – life periods of first order, second order, zero order and third order reactions. Determination of order of reactions. Experimental techniques involved in the following kinetics of reaction. Volumetry, Manometry, Dilatometry, Polarimetry and Colorimetry. Typical examples for each of the techniques.

UNIT-II	**Chemical Kinetics-II**	**12 Hours**

Theoretical aspects: Effect of temperature on the rate constant. The activation energy. Significance of free energy of activation. The collision theory of reaction rates and its limitations. The theory of absolute reaction rates (ARRT). Comparison of the collision theory with the absolute reaction rate theory. Lindemann theory of uni-molecular reactions.

UNIT-III	**Kinetics of Photochemical Reactions**	**12 Hours**

Dark reactions - Complex thermal reactions – Thermal chain reaction - the H2 - Br2 reaction. Absorption of light and photochemical processes – Laws of photochemistry - The Stark–Einstein law of photochemical equivalence. Kinetics of photochemical chain reaction - the H_2 / Br_2 reaction. Quantum yield of photochemical reactions. Comparison of the thermal and photochemical kinetics of the H_2 / Br_2 reaction. Photosensitized reactions – photophysical process, Fluorescence,

Phosphorescence and Chemiluminescence.		
UNIT-IV	**Catalysis and Colloids**	**12 Hours**
***Catalysis – Types of catalysis** – Characteristics of catalytic reactions Theories of catalysis – Intermediate Compound Formation Theory, Adsorption Theory – Acid-base catalysis – Enzyme catalysis – Mechanism of enzyme catalysis – Electrocalatysis. ***Colloids - Definitions** – Classification of colloids – sols – preparation, and properties Stability of colloids, gold number, associated colloids – Cleansing action of soaps and detergents. Emulsion – types of emulsions, preparation and properties. Gels – types of gels, preparation and properties. Applications of colloids.		
UNIT-V	**Liquid State and Liquid Crystals**	**12 Hours**
Intermolecular Forces in Liquids, Dipole-dipole Attractions, London Forces, Hydrogen Bonding Vapour Pressure, Effect of Temperature on Vapour Pressure, Determination of Vapour Pressure The Static Method, The Dynamic Method, Effect of Vapour Pressure on Boiling Points-Surface Tension, Units of Surface Tension, Determination of Surface Tension, Capillary Rise Method, Drop Formation Method, Ring- detachment Method, Bubble Pressure Method – Viscosity, Units of Viscosity, Measurement of Viscosity, Ostwald Method- Effect of Temperature on Viscosity of a Liquid, Refractive Index, Molar Refraction, Determination of Refractive Index Optical Activity, Specific Rotation, Measurement of Optical Activity. Introduction - Liquid crystals, Vapour pressure-temperature diagrams, Thermography, LCDs and the seven segment cell.		
****self study portion***		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Arun Bahl and B. S. Bahl, G. D. Tuli, (2015) Essentials of Physical Chemistry, S. Chand & Co., Revised multicolor edition.	
2.	B. R. Puri, L. R. Sharma, and M. S. Pathania, (2015) Principles of Physical Chemistry, S. Chand & Co., New Delhi.	
3.	P. L. Soni, O. P. Dharmarha and U. N. Dash, (2013) Textbook of Physical Chemistry, S. Chand & Co., New Delhi.	
REFERENCE BOOKS:		
1.	Keith J. Laidler and John H. Meiser,(2014) Physical Chemistry, CBS Publishers & Distributors, New Delhi.	
2.	Gurudeep Raj, (2009) Advanced Physical Chemistry, Goel Publishing House, Meerut.	
3.	K. K. Rohatgi Mukherjee, (2014) Fundamentals of Photochemistry, New age International Publishers.	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	M
CO2	S	S	S	S	S
CO3	S	S	M	H	M
CO4	S	S	H	M	S
CO5	S	S	H	S	H
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Subject code: 23UCH6CN

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PRACTICAL – III INORGANIC QUANTITATIVE ANALYSIS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	VI	3	90	3

Course Objectives

1.	To understand the concept of gravimetric analysis.
2.	To get acquainted with the experimental procedure of gravimetric analysis.
3.	To determine the quantity of analyte in solution.

Course Outcomes (CO)

K1 – K5	CO1	Understand the basic principles of Gravimetric analysis.
	CO2	Understand about the various precipitating agents.
	CO3	Determination of analyte masses through the gravimetric analysis.
	CO4	Improve the accuracy of analysis.
	CO5	To gain knowledge about Metal analysis in cosmetic products using AAS

I	**Gravimetric Analysis**
1	Estimation of Sulphate as Bariumsulphate.
2	Estimation of Barium as Barium Chromate.
3	Estimation of Lead as Lead Chromate.
4	Estimation of Calcium as Calcium Oxalate.
5	Estimation of Nickel as Nickel Dimethylglyoxime.
II.	**Metal analysis in cosmetic products using AAS. (demonstration only)**

TEXT BOOKS :

1.	Venkateswaran. V, Veeraswamy. R, Kulandaivelu . A.R, (1997), **Basic Principles of Practical Chemistry**, New Delhi, Sultan Chand and Sons,

REFERENCE BOOKS:

1.	Mendham. J, Denney, R.C. Bames. J.D and Thomas. (2006) M, **Vogel's Text book of Quantitative Analysis**, Pearson Education.
2.	Gopalan.R, Subramaniam P.S and Rengarajan. (2004) K, **Elements of Analytical Chemistry,** Sultan Chand and Sons.

Time: 3 Hours			**Max. Marks : 50**
Distribution of Total Marks: 50			
	Record	05 marks	
	Gravimetric	45 marks	

Gravimetric Marks-45			
	Procedure (To be written within five minutes)	5 mark	
	Experiment	40 mark	
	Error up to 2%	40 mark	
	3%	35 mark	
	4%	30 mark	
	5%	10 mark	
	>5%	5 mark	

- ❖ Proportionate deduction of marks must be made for errors between the limits given above.
- ❖ Examiners should calculate the result of each candidate with the data obtained by the candidate.
- ❖ For each independent arithmetic error deduct 2 marks.
- ❖ For incomplete or wrong calculation deduct 20% of marks eligible for the result as calculated by the examiners.
- ❖ For incomplete or wrong calculation deduct 30% of marks eligible for the result as calculated by the examiners.
- ❖ If there is a difference between the two results, the examiners should consider the one favorable to the candidate.
  If a candidate is not able to complete the experiment due to accident, award 5 marks

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	H
CO2	S	S	M	H	H
CO3	S	S	H	M	H
CO4	S	S	M	H	H
CO5	S	S	H	H	H

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PRACTICAL – IV PHYSICAL EXPERIMENTS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**VI**	3	90	3

Course Objectives

1.	Transformation of theoretical knowledge gain to practical aspects.
2.	To have experience in handling electrical and non-electrical equipments.
3.	To determine the strength of various solutions through spectrometric and electrochemical techniques.

Course Outcomes (CO)

K1 – K5	CO1	The results of physical chemistry experiments are incorporated in both theoretical and practical aspects.
	CO2	Gain familiarity with a variety of physico-chemical measurement techniques.
	CO3	Interpret data from an experiment, including the construction of appropriate graphs and the evaluation of errors.
	CO4	To know about Determination of Cell Constant, Specific conductivity and Equivalent conductivity of strong electrolyte.
	CO5	To determine strength of acids and bases by Conductometric Titration.

Physical Chemistry Experiments

1.	Determination of Partition coefficient of Iodine between Carbon tetra chloride and water.
2.	Determination of rate constant of acid
3.	Determination of K_f / Molecular weight by Rast Macro method (Naphthalene, Diphenyl and m-dinitrobenzene as solvents).
4.	Determination of Critical solution temperature of Phenol- Water system.
5.	Determination of concentration of an electrolyte (NaCl/ KCl/ Succinic acid)
6.	Determination of Transition temperature of the hydrated salt (Sodium acetate, Sodium thio Sulphate and $SrCl_2.6H_2O$)
7.	Phase diagram-Simple Eutectic system
8.	Determination of Cell Constant, Specific conductivity and Equivalent conductivity of strong electrolyte
9.	Determination of dissociation constant of a weak acid (Acetic acid)
10.	Conductometric Titration (Strong acid Vs Strong base)
11.	Potentiometric Titration (Acid-Base Titration HCl Vs NaOH CH_3COOH Vs Na_2CO_3)
12.	Potentiometric Titration (Redox Titration FAS Vs $KMnO_4$)
13.	Estimation of Manganese by colorimetric method.

TEXT BOOKS :	
1.	Venkateswaran. V, Veeraswamy. R, Kulandaivelu . A.R, (1997), **Basic Principles of Practical Chemistry**, New Delhi, Sultan Chand and Sons,
REFERENCE BOOKS:	
1.	Gopalan. R, Subramaniam. P.S., and Rengarajan, K.,(2004) Elements of Analytica Chemistry, Sultan Chand and Sons**.**

Time: 3 Hours			**Max. Marks : 50**
Distribution of Total Marks: 50			
	Record	05 mark	
	Experiment	45 mark	
Mark Breakup for Experiments : 45			
1. Equilibrium Constant			
For carrying out the experiment	15 mark	Constant error up to 10%	12 mark
Remaining Marks given below	30 mark	10-12%	8 mark
Calculation of [KI]	6 mark	12-14%	4 mark
Calculation of [I_2]	6 mark	>14%	No mark
Calculation of [KI_3] Value of equilibrium	6 mark		
2. Strength of KI Solution may be given in the range from 0.02N to 0.06N			
Calculation of [KI]	12 mark	Calculation	
Calculation of [I_2]	11 mark	Strength of given [KI] Solution	11 mark
Calculation of [KI_3]	11 mark		
	For wrong calculation of above value, 50% of marks to be deducted for those steps only.		
3. HCl or H 2SO4, 0.5N to be given. If the order of difference between theoretical and candidates value is			
Below a factor of 10 Between 10-20 Above - reduce 3 marks for each factor -5 marks		30-45 mark	
4. Critical Solution Temperature of Phenol-Water system is 67.0^{0}C			
Phenol-Water system Plot of % Phenol Vs Temp	45 mark	Error up to ±3°C to 15°C	Reduce 3 marks for each percent
Error up to ±3°C	35 mark	Above ±15°C	5 mark
5. Candidate may be instructed to use solutions of strength 1% and below. Unknown solution			

must be below 0.8%			
Plot of %NaCl Vs Temp	45 mark	Error up to 1-2%	5 mark
Determination of % NaCl Solution	30 mark	Error up to 2-3%	5 mark
Error up to 1%	10 mark	Error above 3%	4 mark
6. 0.1N Acetic acid and 0.1N KCl 100ml each to be given			
Calculation of cell constant	15 mark	Error up to 10%	20 mark
Correct Eq. Conductance of Acetic acid	10 mark	Error up to 10% to 20%	Reduce 1 mark for each %.
7. 0.1N Acetic acid 100ml and 0.1N KCl 100ml to be given			
Calculation of cell constant	20 mark	Calculation of dissociation constant	15 mark
Correct Eq. Conductance of Acetic acid	15 mark	Reduce marks for errors as in **experiment-6 above**	
8. 0.02N HCl may be given			
Error up to 5%	45 mark	Error up to 10-15%	Reduce 3 marks for each %
Error up to 5-10%	Reduce 3 marks for each %	Error above 5%	5 mark
9. Rast method			
Solvent Kf	**Solute**		
1.Naphthalene (6.9°C)	1. Biphenyl	Melting point of solvent	15 mark
2. Diphenyl (8-8.4°C)	2. Naphthalene	Error upto ±10%	30 mark
3. Diphenyl amine (8.4 – 8.8°C)	3. Dicholorobenzene	Error upto ±20%	15 mark
For each step error has to be worked out		Above 20%	5 mark
10. Phase diagram			
Melting point of A	10 mark	Error up to 1%	8 mark
Melting point of B	10 mark	Error up to 2%	6 mark
Construction of phase diagram	9 mark	Error up to 3%	4 mark
Eutectic temperature	8 mark	Error > 3 %	No mark
Eutectic composition	8 mark		
11. Colorimetric Experiments			
Error upto 1%	45 mark	Error upto 3%	20 mark
Error upto 2%	30 mark	Error > 3%	5 mark

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	M	S
CO2	S	S	H	H	S
CO3	S	S	H	H	S
CO4	S	S	S	M	S
CO5	S	S	H	S	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**CORE PRACTICAL – V APPLICATION ORIENTED PRACTICAL**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	VI	4	120	4

Course Objectives

1.	To demonstrate the basic laboratory techniques and application oriented physical constants.
2.	To prepare organic dyes, organic compounds and home care products.
3.	To estimate the hardness of water, DO, available chlorine in bleaching powder and saponification value of an oil.

Course Outcomes (CO)

K1 – K5	CO1	Gain the knowledge of physical constants and preparation of dyes.
	CO2	Know about the preparation of organic compounds.
	CO3	Learn about the preparation method of home care products.
	CO4	Learn about estimation of hardness of water, dissolved oxygen, Saponification of oil and isolation of citric acid.
	CO5	Learn about estimation of hardness of water, dissolved oxygen

I.	**Determination of Physical Constants.**
1	Determination of Melting point
2	Determination of Boiling point
3	Estimation of Lead as Lead Chromate.
4	Estimation of Calcium as Calcium Oxalate.
5	Estimation of Nickel as Nickel Dimethylglyoxime.
II.	**Preparation of Organic dyes**
	Preparation of dyes like Methyl Orange, Methyl Red, Azo Amino benzene.
III	**Preparation of Organic Compounds**
	Preparation involving Acetylation, Hydrolysis, Oxidation, Halogenation, Nitration and Benzoylation.
IV	**Preparation of Home care products**
1	Preparation of white phenyl
2	Preparation of soap oil
3	Preparation of detergent powder
4	Preparation of transparent soap
5	Preparation of moisturizing cream
V	**Estimations**
1	Estimation of Hardness of water using EDTA
2	Estimation of dissolved oxygen in water
3	Estimation of alkalinity in water
4	Estimation of calcium in limestone by EDTA method

5	Estimation of Total Fatty Matter (TFM) of a soap
6	Estimation of acid value of an oil
7	Estimation of available chlorine in bleaching powder
VI	**Day to day activity related experiments.**
1	Natural Product extraction using Rotary vapor evaporator.
2	Water quality analysis using hydrometer.
3	Milk analysis using lactometer.
4	Green synthesis reactions in microwave oven.

TEXT BOOKS :

1.	Venkateswaran. V, Veeraswamy. R, Kulandaivelu . A.R, (1997), **Basic Principles of Practical Chemistry**, New Delhi, Sultan Chand and Sons,
2.	Sharma, K.K. and Sharma, D.S. (2005) **Introduction to Practical Chemistry**, Vikas Publishing House, New Delhi.

REFERENCE BOOKS:

1.	G. Svehla, (1987), **Vogel's Qualitative Inorganic Analysis**, Orient Longman Ltd, Hyderabad.
2.	Praveen Kukreja, (2006). **Chemistry Advanced Practical Manual**, Vrinda Publishing (p) Ltd, New Delhi.

Time: 3 Hours			**Max. Marks : 50**
Distribution of Total Marks: 50			
	Record	05 mark	
	Physical Constant	10 mark	
	Estimation	20 mark	
	Preparation	15 mark	
Physical Constant Marks - (10)			
	±2°	10 mark	
	±3°	8 mark	
	±4°	6 mark	
	±>4°	5 mark	
	Estimation Marks - (20)		
	Error up to 2%	20 mark	
	3%	15 mark	
	4%	10 mark	
	5%	5 mark	
	>5%	3 mark	

	Preparation Marks- (15)		
	Quantity	10 mark	
	Recrystallization/ Quality	05 mark	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	S	S
CO2	S	S	S	S	H
CO3	S	S	S	S	S
CO4	S	S	S	S	H
CO5	S	S	S	H	S
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**MAJOR ELECTIVE - I** **POLYMER TECHNOLOGY**		
Batch		**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024		4	60	5

Course Objectives

1.	To know about basics of polymers, polymerization and plastic materials
2.	To learn about polymer processing and synthesis of some commercially important polymers and to know about various polymer processes techniques.
3.	To know different type of plastics, advancements, disposal, applications

Course Outcomes (CO)

K1 – K5	CO1	Know about the types of polymers, chemical and physical properties, its industrial applications and uses.
	CO2	Understand the various polymerization techniques, processing and different types of individual polymer products.
	CO3	Know about different Polymerization Processing Techniques
	CO4	Acquiring knowledge of commercially important polymer products and its applications.
	CO5	Know about the recent advances in polymer products and their applications.

UNIT – I	**Introduction to polymers**	**12 Hours**

Introduction to polymers: Genesis of polymers: *Basic concept – monomers and polymers – definition. Classification of polymers – natural and synthetic polymers – organic and inorganic Polymers – Thermoplastic and thermosetting plastics – plastic materials, elastomers, few applications.

Chemistry of Polymerization: Types – Chain polymerization – Free radical polymerization, Ionic polymerization, coordination polymerization. Step Polymerization – polycondensation, poly addition, ring opening. Miscellaneous polymerization reactions – electrochemical. Group transfer polymerization.

UNIT-II	**Polymer properties and Reactions**	**12 Hours**

Molecular weight and Size: Average molecular weight - number average - weight average molecular weights – concepts. Sedimentation and viscosity Average molecular weights - Molecular weight and degree of polymerization.

Glass Transition Temperature (Tg): Definition, application of Tg in processing, Transition and Associated properties, Factors influencing the glass transition temperature. Photo oxidative degradation of polymers.

UNIT-III	Polymerization Processing Techniques	12 Hours
Polymer Processing Technology: Introduction to polymer processing-basic concept and applications-Injection Moulding - compression moulding, extrusion process, Blow moulding, Blown film, Rotational Moulding, FRP.		
UNIT-IV	**Chemistry of Commercial Polymers**	**12 Hours**
Chemistry of polymers: Types-engineering and commodity plastics, General methods and preparation Properties of PE, PP, PS, PMMA, polyacetal, PC, polyamides, PVC, PF, UF, epoxy, Teflon.		
UNIT-V	**Recent Advances in Polymers**	**12 Hours**
Polymer and environmental effect-introduction-disposal of polymer waste-recycling system-importance of biopolymers –basis of conducting polymers, introduction to blends and alloys. Application of plastics in various fields-plastic industries in India.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	V. R. Gowariker, N.V. Viswanathan and Jayadev Sreedhar, (2012) **Polymer Science,** Revised Edition, New Age International Publishers, New Delhi.	
REFERENCE BOOKS:		
1.	F.W. Billmeyer – (1990) **Text Book of Polymer Science**, John Wiley & Sons.	
2.	J.R. Fried, (2014) **Polymer Science & Technology**, Prentice Hall of India Private Ltd.	
3.	G.S. Misra, (1997) **Introductory Polymer Chemistry**, New Age International Private Ltd, New Delhi.	
4.	Sharma. B.K., **Polymer Chemistry**, Goel Publishing House, Meerut, (2004).	
5.	Arora. M.G., Singh. M, and Yadav. M.S., (2003) Polymer Chemistry, Anmol Publications Private Ltd., New Delhi.	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	S	S	S
CO2	S	H	S	S	S
CO3	S	S	H	S	S
CO4	S	H	S	H	S
CO5	S	H	S	H	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**MAJOR ELECTIVE - II** **NANO AND GREEN CHEMISTRY**		
Batch		**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024		4	60	3

Course Objectives

1.	To gain knowledge about in - depth look at the basics of Nano Chemistry and to know the methods to prepare Nano materials.
2.	To get the knowledge about Green Chemistry and its limitations.
3.	To have a holistic idea about Green solvents in laboratory as well as in Industry and also to study the Reactions and applications of Green Chemistry.

Course Outcomes (CO)

K1 – K5	CO1	To understand the basics of Nano Chemistry.
	CO2	To know the methods to prepare Nano materials.
	CO3	To have an idea about Nano chemistry in medicine.
	CO4	To gain knowledge about Green reactions in laboratory.
	CO5	To gain knowledge about Green solvents.

UNIT – I	**12 Hours**

Introduction to Nanoscience: Definition of Nanomaterials – classification: Zero Dimensional, One Dimensional, Two Dimensional and Three Dimensional Nanomaterials with an example – Synthesis of Nanomaterials Top Down approach: Photolithography, Conventional Photolithography, Electron Beam lithography and Bottom up Approach: Physical vapor deposition, Plasma Arcing, Laser Ablation, Sol Gel method- Carbon Nanotubes: fullerene, Bukyball 60, Carbon Nanowires – Types, properties and uses.

UNIT-II		**12 Hours**

Preparation of Nanomaterials. Co-precipitation- sol- gel - photochemical reduction – hydrothermal and solvothermal synthesis.

Nano Characterization: Instrumentation – Characterization techniques – SEM, AFM, Powder XRD, Basic Principles, Topography, Morphology, Composition crystalline Structure; Advantages and Disadvantages.

UNIT-III	**12 Hours**

Biologically Inspired Nanotechnology: Basic biological concepts and principles that may lead to the development of technologies for nanoengineering systems.

Nano in Medicine: Drug delivery – Cancer diagnosis & Therapy – In vivo therapy - drug delivery system – Nano biotechnology devices.

UNIT-IV	**12 Hours**

Green Chemistry: Introduction-definition-*Need for green chemistry- Goals - Limitations – Progress of Green Chemistry - principles of green chemistry- Concept of Atom economy- Concept of Selectivity.

Green Reactions: Green reactions-Solvent free synthesis of Aldol condensation (Acid catalyst, Crossed aldol), Claisen rearrangement, Clemmensen reduction, Diels-Alder reaction.

UNIT-V	**12 Hours**

Green Solvents: Green solvents – super critical carbon dioxide, ionic liquids - Water as greener solvent- reactions in ionic-liquid, solvent free reaction. Solvent less reaction – Microwave reactions – sonications.

self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	S. Shanmugam, Nanotechnology, M.J.P. Publishers, Chennai, 2011.
2.	V. Kumar, An Introduction to Green Chemistry, Vishal Publishing Co., 2015.
3.	V.K. Ahluwalia, Green Chemistry, Ane Books India, New Delhi, 2010.

REFERENCE BOOKS:

1.	G. Cao, **Nanostructures & Nano Materials**, synthesis, properties and applications Imperial College Press, U.K, 2004.

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	S	S	S	S
CO2	S	H	S	S	S
CO3	S	S	H	S	S
CO4	S	H	S	H	S
CO5	S	S	S	H	S

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04	B.Sc. Chemistry		
Title of the paper	**MAJOR ELECTIVE - III** **PHARMACEUTICAL CHEMISTRY**		
Batch	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	4	60	5

Course Objectives

1.	To know about the common diseases and cure-terms of pharmacology and drug action.
2.	To get introduced to chemotherapy – antibiotics.
3.	To know the drugs meant for diabetes.

Course Outcomes (CO)

K1 – K5	CO1	Gain the knowledge about the common diseases and cure-terms of pharmacology.
	CO2	Understanding Mechanism of actions of drugs
	CO3	Understand about drug classification.
	CO4	Learn about Common body ailments.
	CO5	Basic ideas about various health promoting drugs.

UNIT – I	**Drugs**	**12 Hours**

Terminology- drug, pharmacology, pharmacognesy, pharmacodynamics, anti metabolites, LD50 and ED50 therapeutic index - Common diseases -infective disease – insect – borne, air borne and water borne - *Various sources of drugs, pharmacologically active constituents in plants, Classification of drugs- biological and chemical classification.

UNIT-II	**Action of Drugs**	**12 Hours**

Mechanism of drug action- Action at cellular and extra cellular sites. Drug receptors and biological responses- Drug receptor interaction through various bonding- Mechanism of different types of drug action- Absorption of drugs- routes of administration of drugs- factors affecting absorption of drug-Assay of drugs-chemical and biological assays.

UNIT-III	**Designation of drugs**	**12 Hours**

Designation of drugs based on physiological action; Definition and two examples each of Anaesthetics-General, IV and local- Definition and two examples each of Analgesics – Narcotic and synthetic- Definition and two examples each of antipyretics and anti-inflammatory agents – Antibiotics –Classification based on biological action, based on chemical structure-Structure and uses of Penicillin, streptomycin, chloramphenicol, tetrecyclines.

UNIT-IV	**Common Body Ailments**	**12 Hours**

Diabetes-types-Causes and control, insulin-Structure and dosage-oral hypoglycemic drugs

(sulphonyl ureas, biguanides)- Blood pressure- hypotension- hypertension (Systolic & Diastolic)- Antihypertensive drugs (Clonidine, alpha methyldopa)- Cardiovascular drugs- therapeutic uses and examples of anti arrhythmic(quinidine), anti anginals (glyceryl trinitrate), vasodilators(hydrallazine hydrochloride)- Antidepressants–types(two examples of each) - hypnotics and sedatives – mechanism of action and therapeutic uses of barbiturates.

UNIT-V	**Health Promoting Drugs**	**12 Hours**

Vitamins A, B, C, D, E and K-Sources, Deficiency diseases, Therapeutic uses - Medicinally important inorganic compounds of Al (aluminium hydroxide gel, alum), P(phosphoric acid), As(arsenous anhydride, sodium arsenate), Hg(yellow mercuric oxide, ammoniated mercury), Fe(ferric ammonium citrate, ferrous sulphate) –preparation- role and applications-Aneamia-symptoms and causes, anti-anemic drug, anticoagulants-Coumarine and Heparin) AIDS- symptoms, prevention, treatment- Cancer-Common causes-treatment –cytotoxic agents-Radioactive isotopes-Hormones and neoplastic agents.

self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	Jayashree Ghosh, (2010) **A Text Book of Pharmaceutical Chemistry**, S.Chand and Company Ltd., New Delhi.

REFERENCE BOOKS:

1.	Lakshmi S., (1995) **Pharmaceutical Chemistry**, S. Chand & Sons, New Delhi.
2.	Ashutosh Kar, (2005) **Medicinal Chemistry**, Wiley Eastern Ltd., New Delhi.
3.	David William & Thomas Lemke, (2018) **Principles of Medicinal Chemistry**, Foyers, BI publishers.
4.	Romas Nogrady, (2008) **Medicinal Chemistry**, Oxford University press 3rd edition.

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	H	H	S	S	S
CO2	S	S	S	S	S
CO3	S	S	S	S	S
CO4	S	S	S	S	S
CO5	S	H	S	H	S

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**MAJOR ELECTIVE - IV** **AGRICULTURAL CHEMISTRY**		
Batch		**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024		4	60	5

Course Objectives

1.	To know about origin, physical and chemical aspects of soil
2.	To know about the basic idea of plant nutrients
3.	To acquire the knowledge pesticides, fungicides and Herbicides

Course Outcomes (CO)

K1 – K5	CO1	To gain the knowledge about the origin soil.
	CO2	To understand about physical and chemical properties of soil.
	CO3	To understand about chemical aspects of soil
	CO4	To learn about plant nutrients.
	CO5	To learn Basic ideas about pesticides, fungicides and herbicides.

UNIT – I	**Origin of soil**	**12 Hours**

Definition of soil-origin-igneous-metamorphic and sedimentary rocks-rock systems-weathering of rocks and minerals - main components of soil-organic, inorganic, liquid and gaseous phase-Minerals of importance with respect to soil, industries and agriculture –Soil formation physical, chemical and biological factors responsible for soil formation-soil forming processes- Core soil groups of Tamilnadu-Soil survey standard soil survey-methods of soil surveys –remote sensing and soil mapping-soil resource management-use of satellite data for source inventory.

UNIT-II	**Physical Properties of Soil**	**12 Hours**

Physical properties of soil-soil texture and textural classification-pore space-bulk density, particle density –soil structure and soil colour-surface area-soil colloids-plasticity, shrinkage-flocculation and deflocculation-soil air, soil temperature, their importance in plant growth-soil reaction –ion exchange reaction-cation exchange-anion exchange –buffering capacity – hydrogen ion concentration-determination of pH Values-factors affecting soil pH-soil pH and nutrient availability- Soil degradation –causes.

UNIT-III	**Chemical Aspects of Soil**	**12 Hours**

Origin of problem soils, their properties acid, alkali and saline soils-diagnosis-remediation of acid and salt effected soils –Methods of reaction and after care-Quality of irrigation water – causes for poor quality waters for irrigation, their effects in soil and crops. Soil testing-Concept,objective and basis-soil sampling, tools, collection processing, dispatch of soil and water samples. Soil organic matter-its decomposition and effect on soil fertility-source of organic matter in soil –maintenance and distribution –soil organism –their role-nitrification-denitrification, nitrogen fixation in soils-

biological nitrogen fixation in soils –microbial interrelationship in soil-microbes in pert and disease management-*__Bio-conversion of agricultural wastes__.

UNIT-IV	**Plant Nutrients**	**12 Hours**

Plant nutrients-macro and micro nutrients-their role in plant growth –sources-forms of nutrient absorbed by plants –factors affecting nutrient absorption-deficiency symptoms in plants-corrective measures-chemicals used for correcting nutritional deficiencies-nutrient requirements of crops, their availability, fixation and release of nutrients. Fertilizers –classification of NPK fertilizers –sources-natural and synthetic –straight –complex –liquid fertilizers, their properties, use and relative efficiency-secondary and micro nutrient fertilizers-mixed fertilizers –principles of fertilizers use – the efficient use of various fertilizers-integrated nutrient management biofertilizers –rhizobium, azospirillum, azotobactor-Blue green algae and azolla production and quality control of bio-fertilizers.

UNIT-V	**Pesticides, Fungicides And Herbicides**	**12 Hours**

Pesticides: Definition –classification –organic and inorganic pesticides-mechanism of action – Characteristics-Safe handling of pesticides –impact of pesticides on soil, plants and environment – Acts and Laws concerning the pesticides-. Fungicides Definition –classification – mechanism of action-Sulphur, copper-mercury compounds, dithanes, dithiocarbamates.

Herbicides: Definition –Classification-mechanism of action-Arsenic and boron compounds-nitro compounds, chloro compounds, Triazines, propionic acid derivatives, urea compounds. Acaricides-Rodenticides-Attractants-Reppellants-Fumiganus foliants.

self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	Biswas ,T.D and Mukeherjee, S.K.(2001) **Textbook of Soil Science**, Tata McGraw – Hill publishing co.
2.	Daji, T.A. (1990) **Textbook of Soil Sciences**, Asia Publishing House, Madras.
3.	Tisdale. S. L., Nelson. W. L. and Beaton. J. D.(2017) **Soil Fertility and Fertilizers**, 8th edition Macmillan Publishing Company, New York.

REFERENCE BOOKS:

1.	Hesse, (1983) **A Textbook of Soil Chemical Analysis** P.R. John Murray.
2.	Buchel, K.H. John Wiley & Sons , (1983) **Chemistry of Pesticides**, New York.
3.	Sree Ramula, (1979) **Chemistry of Insecticides and Fungicides**, U.S. Oxford and IBH Publishing Co., New Delhi.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	H	H	S	S	S
CO2	S	S	S	S	S
CO3	S	S	S	S	S
CO4	S	S	S	S	S
CO5	S	S	H	S	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**MAJOR ELECTIVE -V DAIRY CHEMISTRY**		
Batch		**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024		4	60	5

Course Objectives

1.	To know the chemistry of milk and milk products
2.	To know the basics of milk proteins, milk lipids, milk carbohydrates, and milk vitamins.
3.	To acquire knowledge of dairy products, analyze the constituents of milk products.

Course Outcomes (CO)

K1 – K5	CO1	Learning the chemistry of milk and milk products
	CO2	Knowing the basics of milk proteins, milk lipids, milk carbohydrates, and milk vitamins.
	CO3	Understanding the production and composition of milk products.
	CO4	By applying the acquired knowledge of dairy products, analyze the Constituents of milk products.
	CO5	To know commercial values of milk.

UNIT – I	**12 Hours**

Milk: Definition-General composition of milk-physical properties of milk- colour, odour, acidity-natural and developed, specific gravity-Recknagel effect viscosity and conductivity, factors affecting the gross competition of milk, physico-Chemical change taking place in milk due to processing parameters-boiling pasteurization- sterlilzation and homogenization. Adulterants, preservatives and neutralizers-example and their detection. Estimation of fat, specific gravity, acidy and total solids in milk.

UNIT-II	**12 Hours**

Milk lipids-terminology and definitions classification – saponifiable (triglycerodes) and unsaponifiable matters (sterols and cholesterol) phosphor lipids structure and properties (Lecithin and Cephalin) Milk fat constants-refractive index-saponification number, Iodine number, R.M.number, R.M number and polenske number.

Milk proteins-Chemistry of proteins in general structure-N-terminal and C-terminal, hydrogen bond, disulphide bond and salt linkages, outlines of primary, secondary and tertiary structure of proteins. Physical properties of milk proteins- Electrical properties and hydration, solubility. Reaction of milk properties with formaldehyde and ninhydrin. Non-protein nitrogen constituents of milk, effect of heat on milk protein, milk enzyme and functions.

Milk carbohydrate-Lactose-Its structure, solubility, hydrolysis, Oxidation and reduction, Estimation of lactose in milk.

Milk vitamins-water and soluble vitamins, effect of heat and light on vitamins. Ash and mineral matters in milk.

UNIT-III	**12 Hours**

Creams - Definition-composition-chemistry of creaming process-gravitational and centrifugal methods of separation of cream-Factors influencing cream separation (Mention the factors only)-Cream neutralization. Estimation of fat in cream.

Butter - Definition-% composition-manufacture-Estimation of fat, acidity, salt and moisture content-Desi butter.

Ghee - Major constituents-common adulterants added to ghee and their detection rancidity-definition-types (hydrolytic, oxidative and ketonic) prevention- antioxidants and synergists (natural and synthetic)-Measurements.

UNIT-IV	**12 Hours**

Fermented milk products - Fermentation of milk-definition, conditions, cultured milk-definition of culture-examples, conditions, types-cultured cream-cultured butter milk-Bulgaricus milk-acidophillus milk-yogurt. Racteriophage-definition and its function.

Indigenous products - Definition percentage composition-preparation- physicochemical changes take place during khoa-making-khoa sweet- Gulab jamum, chana sweet-Rossogolla-ingredients and preparation.

Ice cream - Definition-percentage composition-types-ingredients needed manufacture of ice-cream stabilizers-emulsifiers and their role.

UNIT-V	**12 Hours**

Milk powder - Definition-need for making powder-drying process-spray drying, drum drying, jet drying and foam drying-principles involved in each.

Manufacture of whole milk powder by spray drying process-keeping quality of milk powder.

Dairy Detergents - Definition-characteristics-classification-washing procedure (modern method) sterilization-chloramin-T and hypochlorite solution.

self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	**Outlines of Diary Technology**-Sukumar De (2001)
2.	**Principles of Dairy Chemistry**-Robert Jenness & S.Patorn.(1960)
3.	I**ndian Diary products**-K.S. Rangappa and K.T. Achaya. Reprinted (2007)

REFERENCE BOOKS:	
1.	**Modern Diary Products**-L.M. Lampert.
2.	**Principles of Dairy processing** - Warner.

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	H	H	S	S	S
CO2	S	S	S	S	S
CO3	S	S	S	S	S
CO4	S	S	S	S	S
CO5	S	S	S	H	S
S – Strong	**H** – High		**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**MAJOR ELECTIVE - VI LEATHER CHEMISTRY**		
Batch		**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024		4	60	5

Course Objectives

1.	To obtain the knowledge on the structure and composition of the hides, skin and leather.
2.	To know the basic principles involved in the pre-training methods of leather manufacture.
3.	To understand about vegetable tanning, chrome tanning and leather machinery.

Course Outcomes (CO)

K1 – K5	CO1	Learning the basic principles involved in the theory of curing hides and skins
	CO2	Understanding the basics of soaking and bating process
	CO3	Widening a skill on the bating and liming
	CO4	Gaining the broad idea on the Chrome tanning and vegetable tanning process.
	CO5	To know about leather machinery.

UNIT – I	**Theory of curing hides and skins**	**12 Hours**

Putrefaction - Mechanism of Putrefaction – Theory of curing hides and skins- curing by controlling temperature- curing by temperature control- curing by using toxic materials – curing by moisture – Drying – Application of salt – simple salting – brine curing – combined salting and raceway- Indian method of curing.

UNIT-II	**Socking and liming**	**12 Hours**

Socking: Introduction – disinfection of socking water and application of antiseptics- Effect of temperature and pH of sock liquor – Control of socking – surface active agents for leather industry

Liming: Introduction -Effect of liming on collagen, Effect of temperature in liming – neutralization of alkali by collagen –Alkali binding capacity – Shift of isoelectric point due to liming – swelling of collagen.

UNIT-III	**Deliming and bating**	**12 Hours**

Deliming: Introduction – selection of deliming agents- methods of deliming – Extend of deliming – drenching – distinction between deliming and drenching – Chemistry and bacteriology of drenching . Common deliming agents – Boric acid – Ammonium chloride – Ammonium sulphate

Bating : Introduction- objects of bating – important factors in the bating operation : temperature - Duration of bating – strength of bating liquor.

UNIT-IV	Chrome tanning and Vegetable tanning	12 Hours
Chrome tanning: Historical outline – double bath process – Chemical control in double bath process – Firz bath – Dipping – second bath – single bath process – Chemistry of chromium salts – Basic chrome compounds – Olation , polymerization and Oxolation. **Vegetable tanning** : The Chemistry of vegetable tanning materials.		
UNIT-V	**Leather machinery**	**12 Hours**
Introduction – Roller type machine - Unhairing machine – fleshing machine – Scudding machine – Samming and setting machine – Buffing machine – Fluffing wheel – boarding machine – New roller type of machine – Glazing machine – Rolling machine.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	B.K.Sharma, (2019), **Industrial chemistry**, Goel publishing House, Meerut.	
2.	Jayashree Ghosh, Reprint 2008, **Fundamental Concepts of Applied Chemistry**, S.Chand & Company Ltd	
REFERENCE BOOKS:		
1.	P.C.Jain, M.Jain, (2008), **Engineering Chemistry**, Dhanpat Raj Publishing Company Pvt Ltd, 15th Edn.	

MAPPING					
PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	H	H	S	S	S
CO2	S	S	S	S	S
CO3	S	S	S	S	S
CO4	S	S	S	S	S
CO5	S	S	S	H	S
S – Strong	**H** – High		**M** – Medium	**L** – Low	

Programme Code : 04		B.Sc. Chemistry		
Title of the paper		Skill Based Subject 1 – Cyber Security		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	III	2	30	3

Course Objectives

1.	The course introduces the basic concepts of Cyber Security
2.	To develop an ability to understand about various modes of Cyber Crimes and Preventive measures
3.	To understand about the Cyber Legal laws and Punishments

Course Outcomes (CO)

K1 – K5	CO1	To Understand the Concepts of Cybercrime and Cyber Frauds
	CO2	To Know about Cyber Terrorism and its preventive measures
	CO3	To Analyze about the Internet, Mobile Phone and E-commerce security issues
	CO4	To Understand about E-mail and Social Media Issues
	CO5	To Describe about various legal responses to Cybercrime

UNIT – I	**6 Hours**
Introduction to Cyber Security: Definition of Cyber Security- Why is Cyber Security important? Layers of Cyber Security- Evolution of Cyber Security. Cyber hacking - Cyber fraud: Definition- Different modes of cyber fraud - Cyber fraud in India. Cyber pornography.	
UNIT-II	**6 Hours**
Cyber Terrorism: Modes of cyber terrorism. Cybercrime: What is Cybercrime? Cybercrime preventive methods - Preventive steps for individuals & organizations - Kinds of cybercrime - Malware and its types – Cyber attacks.	
UNIT-III	**5 Hours**
Internet Mobile Phone and E-commerce Security issues: Data theft - Punishment of data theft- Theft of internet hours - Internet safety tips for children & parents. Mobile phone privacy- E-Commerce security issues.	
UNIT-IV	**6 Hours**
Email and Social media issues: Aspects of Social Media - The Vicious Cycle of unhealthy social media use- Modifying social media use to improve mental health. Computer Virus - Antivirus – Firewalls.	
UNIT-V	**7 Hours**
Cyber Forensics and Digital Evidence: What does Digital Footprint Mean? - Web Browsing and Digital Footprints- Digital Footprint examples – How to Protect Your Digital Footprints? - How to erase your Footprints? - Browser Extensions and Search Engine Deletion - Cyber Crime and Cyber	

Laws - Common Cyber Crimes and Applicable Legal Provisions: A Snapshot - Cyber Law (IT Law) in India – The Information Technology Act of India 2000 - Cyber Law and Punishments in India - Cyber Crime Prevention guide to users – Regulatory Authorities.

TEXT BOOKS:	
1.	**"Cyber Security", Text Book prepared by** "Kongunadu Arts and Science College", Coimbatore -29, 2022.
REFERENCE BOOKS:	
1.	**Mayank Bhushan, Rajkumar Singh Rathore, Aatif Jamshed, "Fundamental of Cyber Security", BPB Publications, 1st Edition, 2017.**
2.	Anand Shinde, "**Introduction to Cyber Security-Guide to the world of Cyber Security**",Notion Press,2021.
3.	Paul Grishman, "**Cyber Terrorism- The use of the Internet for Terrorist Purpose**",Axis Publication,1st Edition 2010.
4.	Shilpa Bhatnagar, "**Encyclopaedia of Cyber and Computer Hacking**", Anmol Publications, 1st Edition 2009.

Web References:	
1.	http://deity.gov.in/ - Department of Electronics and Information Technology,
2.	Govt. of India
3.	http://cybercellmumbai.gov.in/ - Cybercrime investigation cell
4.	http://ncrb.gov.in/ - National Crime Records Bureau
5.	http://catindia.gov.in/Default.aspx - Cyber Appellate Tribunal
6.	http://www.cert-in.org.in/ - Indian Computer Emergency Response Team
7.	http://cca.gov.in/rw/pages/index.en.do - Controller of Certifying Authorities
8.	www.safescrypt.com - Safescrypt
9.	www.nic.in – National Informatics Centre
10.	https://www.kaspersky.com/resource-center/definitions/what-is-a-digital-footprint
11.	https://geekflare.com/digital-footprint/

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**SKILL BASED SUBJECT-II** **WATER POLLUTION AND MANAGEMENT**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**IV**	2	30	3

Course Objectives

s1.	To know about the sources and characteristics of water.
2.	To learn about the analysis of the pollutants in water.
3.	To learn the methods of purification and management of water.

Course Outcomes (CO)

K1 – K5	CO1	To understand the importance of water.
	CO2	To study the types of water pollution.
	CO3	To analyze and measure the toxic chemical substances.
	CO4	To gain knowledge in purification techniques of water.
	CO5	To know the irrigation systems used in agriculture.

UNIT – I	**Sources of water and its importance**	**6 Hours**

Introduction- Sources of water – Chemistry of water – Water quality parameter- Domestic purpose - Types of water- hard and soft water- uses of water – disadvantages of using hard water for domestic purpose- disadvantages of using hard water for various industries - water for industry. Quality of natural waters - Chemistry of water - Water in human body - Water as a solvent - main quality characteristics of water - effects of water on rocks and minerals - organic matter in water - humic material in water - colloidal material in water.

UNIT-II	**Types and sources of water pollution**	**6 Hours**

*Definition - water pollutants - Sewage and domestic wastes- Industrial effluents- Agricultural discharges- Fertilizers- Detergents- Toxic metals- Siltation- Thermal pollutants- Radioactive materials in water- types of water pollution - ground water pollution, surface water pollution, lake water pollution, river water pollution and sea water pollution - physical pollution of water - chemical pollution of water - biological pollution of water – physiological pollution of water.

UNIT-III	**Analysis of a water pollutants**	**6 Hours**

Physical and chemical examination of water- Sample - preservation and pre concentration method- carbon adsorption method, freeze concentration method, solvent extraction method. Chemical substances affecting portability - suspended solids - dissolved solids - alkalinity - measurement of toxic chemical substances – Radioactivity of water-methods of removing radioactivity from water. Dissolved oxygen - Biochemical Oxygen Demand - Chemical Oxygen Demand.

UNIT-IV	**Purification of water**	**6 Hours**
Removal of coarse, dispersed and colloidal impurities from water - Coagulation of water - Contact and electrochemical coagulation - Flocculants - Purification or treatment of water for municipal supply (Screening - Aeration - Sedimentation with coagulation- Filtration- Sterilisation and disinfection – Storage and distribution) - Chemical methods of sterilisation - Physical methods of sterilization. Desalination of brackish water - Reverse osmosis. Zeolite process - Ion exchange method - Demineralization of water.		
UNIT-V	**Water management**	**6 Hours**
Introduction - Water management - Use and conservation of water resources - Water quality management - Rain water harvesting - Water from rocks - Water management in agriculture - Rain fed system - Irrigated systems - Sea water for agriculture - Water management in industries.		
Self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	B.K Sharma, (2017) **Environmental Chemistry,** Goel publishing House, Meerut.	
2.	A. K. Dee ,(2017) **Environmental chemistry**, New age international.	
REFERENCE BOOKS:		
1.	Jain and Jain, (2001) **Engineering Chemistry**, Dhanpat Rai Publishing Co.,	
2.	N. Manivasakam (Water Analyst),(2001) **Chemical and Microbiological Analysis of Mineral Water and Packaged Drinking Water,** Principal Public Health Laboratory, Coimbatore.	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	H	M	H	M	S
CO2	M	S	S	S	S
CO3	H	S	M	S	S
CO4	M	S	S	M	S
CO5	H	S	S	H	H
S – Strong	**H** – High		**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**SKILL BASED SUBJECT – III FOOD CHEMISTRY**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	VI	2	30	3

Course Objectives

1.	To have an idea about food adulteration and food preservation techniques.
2.	To understand the chemistry of vinegar, fruit juices, vegetable acids and beverages.
3.	To analyse and characterize chemical aspects of milk.

Course Outcomes (CO)

K1 – K5	CO1	Know about the nutrition values in food, food adulteration, standards of food, contamination and food poisoning.
	CO2	Understand about the minerals in food
	CO3	Know about food additives
	CO4	Understand the detailed information about commercially important beverages
	CO5	Know about dairy products

UNIT – I	**Nutrition values of food materials and food adultration**	**6 Hours**

Nutritional values of carbohydrates, vitamins, proteins and fats. ***Definition – Food standards** – **Food Standards in India** – Standards for ensuring quality of Products – Common adulterants in different foods – Contamination of foods with toxic chemicals, pesticides and insecticides. Contamination of food with harmful micro-organisms – Bacterial infections – Fungal contaminations of foods.

UNIT-II	**Minerals in food**	**6 Hours**

Introduction – Minerals in Food - Calcium – Phosphorus - Magnesium – Zinc - Copper – Fluoride – Iodine - Iron – Sodium – Chloride - Potassium Functions - Deficiency problems - Food sources.

UNIT-III	**Food additives**	**6 Hours**

Food Additive- Direct additives / intentional additives - Indirect Additives / Unintentional- Functions of Intentional Food Additives. Major Food Additive Categories- Color Additives- Lakes - Flavor Additives- The Sensation Of Taste - Tripartite Model. Artificial Sweeteners – Aspartame – Cyclamate – Saccharin. Acesulfame Potassium (Acesulfame K). Flavor Intensifier/Flavor Enhancer (Potentiator) - Monosodium Glutamate (Msg). Color And Flavor Preservatives – Antioxidants - Chelating agents. Fat Substitutes – Olestra – Polydextrose – Simplesse. Nutritional Additives - Texture-Improving Additives - Anticaking Agents - Dough Conditioners by pH Value, Alcohol in Vinegar. Fruit Juices and Vegetable Acids: Examination of Lime Juice, Lemon Squash, etc.,

UNIT-IV	**Commercially important beverages**	**6 Hours**
Introduction – Types of beverages - Tea - Nature and Properties of Tea - Adulteration of Tea - Tea Infusions. Coffee - Nature and Properties of Coffee - Adulteration of Coffee with Chicory. Cocoa and Chocolate - Nature and properties of Cocoa and Chocolate - Adulteration of Cocoa. Alcoholic Beverages - Introduction - Determination of Alcohol - Proof Spirit - Denaturing of Alcohol.		
UNIT-V	**Dairy products**	**6 Hours**
General characteristics of Cows' milk –The determination of the specific gravity – use of lactometer – Determination of total solids – Determination of ash – Determination of fat [Gerber, gottlieb – rose, Werner – Schmid and Adams methods] – Calculation of extent of adultration - Determination of total protein – Determination of lactose – Determination of acitidy – Added colouring matter in milk – Preservatives in milk [Boric acid, formaldehyde, etc.] – Boiled and pasturised milk – Homogenised milk – Cream – Reconstituted cream – Synthetic cream – Condensed milk – Dried milk.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		

1.	C. Kenneth Tinkler and Helen Masters,(2005) **Applied Chemistry**, Vol. II, London.
2.	M. Swaminathan, (2003) **Advanced Text Book on Food and Nutrition**, Vol. II, .
3.	B. Sri Lakshmi, (2002) **Nutrition science**, New Age International Pvt. Ltd., New Delhi.

REFERENCE BOOKS:

1.	M. Swaminathan, (2004) **Handbook of Food and Nutrition**, The Bangalore Printing and Publishing Co. Ltd.,
2.	B. Sri Lakshmi, (2004) **Food Science**, New Age International Pvt. Ltd., New Delhi.
3.	S. D. Venkataiah,(2004) **Nutrition Education**, Anmol Publication Pvt. Ltd.,.

MAPPING

PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	H	S	S	S	S
CO2	H	H	S	S	S
CO3	H	S	H	S	S
CO4	H	H	S	S	S
CO5	H	H	H	S	S

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**EXTRA DEPARTMENTAL COURSE (EDC) - CHEMISTRY IN DAY TODAY LIFE**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	V	2	30	3

Course Objectives

1.	To gain knowledge about water treatment in industrial plant and its usage.
2.	To get the knowledge about industrial fermentation process, oil, wax and soap preparation.
3.	To have a holistic idea about food adulteration, food hygiene and paints manufacture.

Course Outcomes (CO)

K1 – K5	CO1	Basic understanding of water technology and acquire knowledge in the treatment of water for multi-purpose.
	CO2	To study about Vitamins in Food
	CO3	To understand the chemistry involved in the manufacturing process of oil, fats, wax and soap.
	CO4	To design a demonstration, that provides an opportunity to identify adulteration in food standards.
	CO5	Broadening the knowledge about paints and pigments.

UNIT – I	**Water Treatment**	**6 Hours**

Introduction – Sources and Uses of Water – Water for Industrial Purposes – Quality of Normal water – water in human body – Hardness of water – Types - Softening of Water – Soda Lime Process, Zeolite, and Ion-exchange Processes (principles only). Demineralization of water – Treatment of Water for Municipal purposes – Desalination of Brackish Water – Electro dialysis – Reverse Osmosis Method (principles only).

UNIT-II	**Vitamins in Foods**	**6 Hours**

Introduction - Fat Soluble Vitamins - Vitamin A - Retinol activity equivalency (RAE) – Immunity - Growth and development - Red blood cell production - Stability of vitamin A. Vitamin D - Importance of vitamin D - Calcium Balance - Insulin secretion - Blood pressure regulation. Water-Soluble Vitamins – Thiamin (Vitamin B1) - Importance of Thiamin - Riboflavin (Vitamin B2) - Importance Of Riboflavin - Niacin (Vitamin B3) - Importance of Niacin.

UNIT-III	**Oils, Fats, Waxes and Soap**	**6 Hours**

Waxes – Classification – Solubility – Saponification value –Manufacture of Candles – Hydrocarbon of Candles – Hydrogenation – of Oils – Soaps –Manufacture – detergents – Cleansing Action of Soaps.

UNIT-IV	Food Adulteration and Hygiene	6 Hours
Definition of Adulteration Food – Common Adulterants in Different Foods – Toxic Effects of Some Metals and Chemicals – Contamination of Foods with Harmful Microorganisms – **Detection of Adulteration in Some Common Food items – *Food Additives and Preservatives** – Food standards.		
UNIT-V	**Paints**	**6 Hours**
Classification – Requirements of a Good Paint and Importance of pigment volume concentration (PVC) – Paints Failure. Emulsion Paints , Enamels , Lacquers and Varnishes – constituents and Manufacture.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	B.K. Sharma, **Environmental Chemistry**, Krishna Prakasam Medai (P) Ltd., Meerut, 6th Revised Edn., (2011).	
2.	P.C. Jain and Monika Jain, **Engineering Chemistry**, Dhanpat Rai & Sons, Delhi, 16th edition (2019)	
3.	M. Swaminathan, **Food & Nutrition, Bappco**, 2nd ed. (2011).	
REFERENCE BOOKS:		
1.	B. Sri Lakshmi, Food Science, New Age, 5th ed. (2013).	
2.	Jayashree, Applied Chemistry, S. Chand, 3rd ed. (2013).	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	H	S	S	S	S
CO2	H	S	S	S	S
CO3	M	H	S	S	S
CO4	H	H	S	S	S
CO5	H	H	S	S	S
S – Strong	**H** – High		**M** – Medium	**L** – Low	

Programme code : 04				
Title of the paper		**PART IV – ENVIRONMENTAL STUDIES**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	I	2	**30**	**2**

Course Objectives

1.	The course will provide students with an understanding and appreciation of the complex interactions of man, health and the environment. It will expose students to the multi-disciplinary nature of environmental health sciences
2.	To inculcate knowledge and create awareness about ecological and environmental concepts, issues and solutions to environmental problems.
3.	To shape students into good "Ecocitizens" thereby catering to global environmental needs.
4.	This course is designed to study about the types of pollutants including gases, chemicals petroleum, noise, light, global warming and radiation as well as pollutant flow and recycling and principles of environmental pollution such as air, water and soil
5.	The course will address environmental stress and pollution, their sources in natural and workplace environments, their modes of transport and transformation, their ecological and public health effects, and existing methods for environmental disease prevention and remediation

Course Outcomes (CO)

K1 – K5	CO1	Understand how interactions between organisms and their environments drive the dynamics of individuals, populations, communities and ecosystems.
	CO2	Develop an in depth knowledge on the interdisciplinary relationship of cultural, ethical and social aspects of global environmental issues.
	CO3	Acquiring values and attitudes towards complex environmental socio-economic challenges and providing participatory role in solving current environmental problems and preventing the future ones.
	CO4	To gain inherent knowledge on basic concepts of biodiversity in an ecological context and about the current threats of biodiversity.
	CO5	To appraise the major concepts and terminology in the field of environmental pollutants, its interconnections and direct damage to the wildlife, in addition to human communities and ecosystems.

UNIT – I	**MULTIDISCIPLINARY NATURE OF ENVIRON -MENT**	**6 Hours**

Definition: scope and importance – Need for public awareness - Natural resources – Types of

resources – Forest Resources – Water Resources – Mineral Resources – Food Resources – Energy Resources – Land Resources.

UNIT-II	**ECOSYSTEMS**	**6 Hours**

Concept of an ecosystem – Structure and functions of an ecosystem – Procedures, consumers and decomposers – Energy flow in the ecosystem – Ecological succession – Food chains, food web and ecological pyramids – Structure and function of the following ecosystem – Forest Ecosystem –Grassland Ecosystem – Desert Ecosystem – Aquatic Ecosystem.

UNIT-III	**BIODIVERSITY AND ITS CONSERVATION**	**6 Hours**

Introduction – Definition – Genetic – Species and ecosystem diversity- Bio geographical classification of India – Value of biodiversity – Biodiversity at global, national and local levels – India as a mega - diversity Nation - Hot spot of biodiversity – Threats to biodiversity - Endangered and endemic species of India – Conservation of Biodiversity – insitu Conservation of Biodiversity – *Exsitu* Conservation of Biodiversity.

UNIT-IV	**ENVIRONMENTAL POLLUTION**	**6 Hours**

Definition - Causes, effects and control measures of: Air Pollution – Water Pollution – Soil Pollution – Marine Pollution – Noise Pollution – Thermal Pollution – Nuclear Pollution – Solid Waste Management: Causes, effects, control measures of urban and industrial wastes – Role of individual in prevention of pollution – Pollution case studies – domestic waste water, effluent from paper mill and dyeing, cement pollution – Disaster Management – Food, Drought, Earthquake, Tsunami, Cyclone and Landslide.

UNIT-V	**SOCIAL ISSUES AND THE ENVIRONMENT**	**6 Hours**

Sustainable Development – Urban problems related to energy – Water Conservation: Rain Water Harvesting and Watershed Management – Resettlement and rehabilitation of people, its problems and concerns, case studies Narmatha Valley Project – Environmental ethics, issues and possible solutions – Climate change, global warming, ozone layer depletion, acid rain, nuclear accidents and holocaust, case studies – Hiroshima and Nagasaki, Chernobyl – Consumerism and waste products – Environmental Protection Act – Air Pollution Act (Prevention and Control) – Water Pollution Act (Prevention and control) – Wild Life Protection Act – Forest Conservation Act – Issues involved in enforcement of environmental legislation – Public awareness – Human Population and the environment – Population Growth and Distribution – Population Explosion – Family Welfare Programme – Environment and Human Health – Human Rights – Value Education – HIV/ AIDS – Women and Child Welfare – Role of Information Technology in Environment and Human Health.

Teaching Methods: Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS:

1.	P.Arul, A Text Book of Environmental Studies, Environmental Agency, No 27, Nattar street, Velacherry main road, Velacheery, Chennai – 42, First Edition, Nov.2004.

REFERENCE BOOKS:

1.	Purohit Shammi Agarwal, A text Book of Environmental Sciences, Publisher Mrs.Saraswati Prohit, Student Education , Behind Naswan Cinema Chopansi Road, Jodhpur.
2.	Dr.Suresh and K.Dhameja, Environmental Sciences and Engineering, Publisher S.K.Kataria & Sons, 424/6, Guru Nanak Street, Vaisarak, Delhi -110 006.
3.	J.Glynn Henry and Gary W Heinke, Environmental Science and Engineering, Prentice Hall of India Private Ltd., New Delhi – 110 001.

Question Paper Pattern **(External only)**	
Duration: 3 Hours	**Total Marks: 50**
Answer all Questions (5 x 10 = 50 Marks)	
Essay type, either or type questions from each unit.	

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**MORAL AND ETHICS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	II	2	30	2

Course Objectives

1.	To impart Value Education in every walk of life.
2.	To help the students to reach excellence and reap success.
3.	To impart the right attitude by practicing self-introspection.
4.	To portray the life and messages of Great Leaders.
5.	To insist the need for universal brotherhood, patience and tolerance.
6.	To help the students to keep them fit.
7.	To educate the importance of Yoga and Meditation.

Course Outcomes (CO)

K1 – K5	CO1	Will be able to recognize Moral values, Ethics, contribution of leaders, Yoga and its practice
	CO2	Will be able to differentiate and relate the day to day applications of Yoga and Ethics in real life situations
	CO3	Can emulate the principled life of great warriors and take it forward as a message to self and the society
	CO4	Will be able to Analyse the Practical outcome of practicing Moral values in real life situation
	CO5	Could Evaluate and Rank the outcome of the pragmatic approach to further develop the skills

UNIT – I	**Moral and Ethics**	**4 Hours**

Introduction – Meaning of Moral and Ethics – Social Ethics – Ethics and Culture – Aim of Education.

UNIT-II	**Life and Teachings of Swami Vivekananda**	**6 Hours**

Birth and Childhood days of Swami Vivekananda – At the Parliament of Religions – Teachings of Swami Vivekananda.

UNIT-III	**Warriors of our Nation**	**4 Hours**

Subhas Chandra Bose – Sardhar Vallabhbhai Patel – Udham Singh – V. O. Chidambaram Pillai – Bhagat Singh – Tiruppur Kumaran – Dheeran Chinnamalai – Thillaiaadi Valliammai – Velu Nachiyar – Vanchinathan.

UNIT-IV	**Physical Fitness and Mental Harmony**	**8 Hours**
Simplified Physical Exercise – Hand Exercises – Leg Exercises – Neuro Muscular Breathing Exercises – Eye Exercises – Kabalabathi – Maharasana A & B – Massage - Acupressure – Relaxation – Kayakalpa Yogam - LifeForce – Aim & Objectives – Principle – Methods. Introspection – Analysis of Thoughts – Moralization of Desires – Neutralization of Anger – Eradication of Worries		
UNIT-V	**Yoga and Meditation – The Asset of India**	**8 Hours**
Yogasanam – Rules & Regulations – Surya Namaskar – Asanas –Sitting – Stanging – Prone - Supine - Pranayama – Naadi Sudhi – Ujjayi – Seethali – Sithkari - Benefits. Meditation – Thanduvasudhi - Agna – Shanthi – Thuriyam – Benefits.		
***Self study portion**		
Teaching Methods: Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class.		
TEXT BOOKS:		
1.	Value Based Education – **Moral and Ethics** – Published by Kongunadu Arts and Science College (Autonomous), Second Edition (2021).	
REFERENCE BOOKS:		
1.	**Swami Vivekananda – A Biography**, Swami Nikhilananda, Advaita Ashrama, India, 24th Reprint Edition (2010).	
2.	**Gandhi, Nehru, Tagore and other eminent personalities of Modern India**, Kalpana Rajaram, Spectrum Books Pvt. Ltd., revised and enlarged edition(2004).	
3.	**Freedom Fighters of India**, Lion M.G. Agrawal, Isha Books Publisher, First Edition (2008).	
4.	**Easy steps to Yoga by Swami Vivekananda**, A Divine Life Society Publication (2000).	
5.	**Yoga Practices - 1** – The World Community Service Centre – Vethathiri Publications Sixth Edition (2017), Erode.	
6.	**Yoga Practices - 2** – The World Community Service Centre – Vethathiri Publications – Eighth Edition (2017), Erode.	
Question Paper Pattern **(External only)**		

Subject code: 23UHR3N1

Duration: 3 hours			**Total Marks**: 50	
Answer all Questions (**5 x 10 = 50 Marks**)				
Essay type, either or type questions from each unit.				
Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**PART IV -NON MAJOR ELECTIVE –I HUMAN RIGHTS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	III	2	30	2

Course Objectives

1.	To prepare for responsible citizenship with awareness of the relationship between Human Rights, democracy and development.
2.	To impart education on national and international regime on Human Rights.
3.	To sensitive students to human suffering and promotion of human life with dignity.
4.	To develop skills on human rights advocacy.
5.	To appreciate the relationship between rights and duties.
6.	To foster respect for tolerance and compassion for all living creature.

Course Outcomes (CO)

K1 – K5	CO1	To understand the hidden truth of Human Rights by studying various theories.
	CO2	To acquire overall knowledge regarding Human Rights given by United Nation Commission. (UNO)
	CO3	To gain knowledge about various organs responsible for Human Rights such as National Human Rights Commission and state Human Right commission (UNHCR).
	CO4	To get habits of how to treat aged person, others and positive social responsibilities.
	CO5	To treat and confirm, child, refugees and minorities with positive social justice.

UNIT – I	**6 Hours**
Definition, Meaning, Concept ,Theories and Kinds of Human Rights- Evaluation and Protection of Human Rights in India- Development of Human Rights under the United Nations.	

UNIT-II	**6 Hours**
United Nations Charter and Human Rights - U.N.Commission on Human Rights- Universal Declaration of Human Rights - International Covenant on • Civil & Political Rights • Economic, Social and Cultural Rights	

UNIT-III	**6 Hours**

Human Rights and Fundamental Rights (Constitution) - Enactments regarding Human Rights Laws in India - National Human Rights Commission and State Human Rights Commission.	
UNIT-IV	**6 Hours**
Aged persons and their Human Rights - Human Rights of Persons with Disabilities - Tribal Human Rights in India - Three Generation Human Rights -Social Awareness and Responsibilities of Individuals.	
UNIT-V	**6 Hours**
Rights of Women, Child, Refugees and Minorities –Social media and Human Rights - NGO's in protection of Human Rights - Right to Election.	
self study portion	
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class	
TEXT BOOKS :	
1.	**Human Rights** (2019) Published by Kongunadu Arts and Science College, Coimbatore –29.
REFERENCE BOOKS:	
1.	**Human Rights**, (2018), Jaganathan,MA.,MBA.,MMM.,ML.,ML., Humanitarian Law and Refugee Law, J.P.Arjun Proprietor,Usha Jaganathan, law series, 1st floor, Narmatha Nanthi Street, Magathma Gandhi Nagar, Madurai – 625014.

Question Paper Pattern (External only)	
Duration: 3 hrs	Max: 75 marks
Section A (5x5=25)	
Short notes Either – Or/ Type - Question from each unit	
Section B (5X10=50)	
Essay type Either – Or/ Type - Question from each unit	

Subject code : 23UWR4N2

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**Part IV -NON- MAJOR ELECTIVE – II WOMEN'S RIGHTS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	IV	2	30	2

Course Objectives

1.	To know about the laws enacted to protect Women against violence.
2.	To impart awareness about the hurdles faced by Women.
3.	To develop a knowledge about the status of all forms of Women to access to justice.
4.	To create awareness about Women's rights.
5.	To know about laws and norms pertaining to protection of Women.
6.	To understand the articles this enables the Women's rights.
7.	To understand the Special Women Welfare laws.
8.	To realize how the violence against Women puts an undue burden on healthcare services.

Course Outcomes (CO)

K1 – K5	CO1	Understand the importance of Women's Studies and incorporate Women's Studies with other fields.
	CO2	Analyze the realities of Women Empowerment, Portrayal of Women in Media, Development and Communication.
	CO3	Interpret the laws pertaining to violence against Women and legal consequences.
	CO4	Study the important elements in the Indian Constitution, Indian Laws for Protection of Women.
	CO5	To be Aware of Government Developmental schemes for women and to create Awareness on modernization and impact of technology on Women.

UNIT – I	**Women's Studies:**	**6 Hours**

Basic concepts of Women's studies in Higher education, Women's studies perspectives- Socialization- Patriarchy- Women's studies as an academic discipline- Growth and development of Women's studies as a discipline internationally and in India.

UNIT-II	**Socio-Economic Development of Women**	**6 Hours**

Family welfare measures, role of Women in economic development, representation of Women in media, status of Women land rights, Women Entrepreneurs, National policy for the empowerment of women.

UNIT-III	**Women's Rights – Access to Justice**	**6 Hours**

Crime against Women, domestic violence – physical abuse- verbal abuse – emotional abuse - economic abuse – minorities, dowry- harassment and death, code of conduct for work place, abetment of suicide.		
UNIT-IV	**Women Protective acts**	**6 Hours**
Protective legislation for Women in the Indian constitution- Anti dowry, SITA, PNDT, and Prevention Sexual Harassment at Workplace (Visaka case), Domestic violence (Prevention) Act.		
UNIT-V	**Women and Child welfare**	**6 Hours**
Safety provisions - various forms of mass media, radio, visual, internet, cyber space, texting, SMS and smart phone usage. Healing measures for the affected Women and child society by private and public sector, NGO and society.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	**Women's Rights** (2021), Published by Kongunadu Arts & Science College, Coimbatore – 641 029.	
REFERENCE BOOKS:		
1.	"**Rights of Indian Women**" by Vipul Srivatsava. Publisher: Corporate Law Advisor, 2014.	
2.	"**Women's security and Indian law**" by Harsharam Singh. Publisher: Aabha Publishers and Distributors, 2015.	
3.	"**Women's Property Rights in India"** by Kalpaz publications, 2016.	

Question Paper Pattern (External only)	
Duration: 3 hrs	**Max:** 75 marks
Section A (5x5=25)	
Short notes Either – Or/ Type - Question from each unit	
Section B (5X10=50)	
Essay type Either – Or/ Type - Question from each unit	

Programme Code : 04	B.Sc. Chemistry		
Title of the paper	**NON- MAJOR ELECTIVE – CONSUMER AFFAIRS**		
Batch	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	2	30	2

Course Objectives

1.	To familiarize the students with their rights and responsibilities as a consumer.
2.	To understand the procedure of redress of consumer complaints.
3.	To know more about decisions on Leading Cases by Consumer Protection Act.
4.	To get more knowledge about Organizational set-up under the Consumer Protection Act
5.	To impart awareness about the Role of Industry Regulators in Consumer Protection
6.	To understand Contemporary Issues in Consumer Affairs

Course Outcomes (CO)

K1 – K5	CO1	Able to know the rights and responsibility of consumers.
	CO2	Understand the importance and benefits of Consumer Protection Act.
	CO3	Applying the role of different agencies in establishing product and service standards.
	CO4	Analyse to handle the business firms' interface with consumers.
	CO5	Assess Quality and Standardization of consumer affairs

UNIT – I	**6 Hours**

Conceptual Framework - Consumer and Markets: Concept of Consumer, Nature of markets: Liberalization and Globalization of markets with special reference to Indian Consumer Markets, E-Commerce with reference to Indian Market, Concept of Price in Retail and Wholesale, Maximum Retail Price (MRP), Fair Price, GST, labeling and packaging along with relevant laws, Legal Metrology. Experiencing and Voicing Dissatisfaction: Consumer buying process, Consumer Satisfaction/dissatisfaction-Grievances-complaint, Consumer Complaining Behaviour: Alternatives available to Dissatisfied Consumers; Complaint Handling Process: ISO 10000suite.

UNIT-II	**6 Hours**

The Consumer Protection Law in India - Objectives and Basic Concepts: Consumer rights and UN Guidelines on consumer protection, Consumer goods, defect in goods, spurious goods and services, service, deficiency in service, unfair trade practice, restrictive trade practice.

Organizational set-up under the Consumer Protection Act: Advisory Bodies: Consumer Protection Councils at the Central, State and District Levels; Adjudicatory Bodies: District Forums, State Commissions, National Commission: Their Composition, Powers, and Jurisdiction (Pecuniary and Territorial), Role of Supreme Court under the CPA with important case law.

UNIT-III	**6 Hours**
Grievance Redressal Mechanism under the Indian Consumer Protection Law - Who can file a complaint? Grounds of filing a complaint; Limitation period; Procedure for filing and hearing of a complaint; Disposal of cases, Relief/Remedy available; Temporary Injunction, Enforcement of order, Appeal, frivolous and vexatious complaints; Offences and penalties. Leading Cases decided under Consumer Protection law by Supreme Court/National Commission: Medical Negligence; Banking; Insurance; Housing & Real Estate; Electricity and Telecom Services; Education; Defective Products; Unfair Trade Practices.	
UNIT-IV	**6 Hours**
Role of Industry Regulators in Consumer Protection i. Banking: RBI and Banking Ombudsman ii. Insurance: IRDA and Insurance Ombudsman iii. Telecommunication: TRAI iv. Food Products: FSSAI v. Electricity Supply: Electricity Regulatory Commission vi. Real Estate Regulatory Authority	
UNIT-V	**6 Hours**
Contemporary Issues in Consumer Affairs - Consumer Movement in India: Evolution of Consumer Movement in India, Formation of consumer organizations and their role in consumer protection, Misleading Advertisements and sustainable consumption, National Consumer Helpline, Comparative Product testing, Sustainable consumption and energy ratings. Quality and Standardization: Voluntary and Mandatory standards; Role of BIS, Indian Standards Mark (ISI), Ag-mark, Hallmarking, Licensing and Surveillance; Role of International Standards: ISO an Overview. Note: Unit 2 and 3 refers to the Consumer Protection Act, 2086. Any change in law would be added appropriately after the new law is notified.	
self study portion	
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class	
SUGGESTED READINGS:	
1.	Khanna, Sri Ram, Savita Hanspal, Sheetal Kapoor, and H.K. Awasthi. (2007) Consumer Affairs, Universities Press.
2.	Choudhary, Ram Naresh Prasad (2005). Consumer Protection Law Provisions and Procedure, Deep and Deep Publications PvtLtd.
3.	G. Ganesan and M. Sumathy. (2012). Globalisation and Consumerism: Issues and Challenges,

	Regal Publications.
4.	Suresh Misra and Sapna Chadah (2012). Consumer Protection in India: Issues and Concerns, IIPA, New Delhi.
5.	Rajyalaxmi Rao (2012), Consumer is King, Universal Law Publishing Company.
6.	Girimaji, Pushpa (2002). Consumer Right for Everyone Penguin Books.
7.	E-books :-www.consumereducation.in
8.	Empowering Consumers e-book,www.consumeraffairs.nic.in
9.	ebook,www.bis.org
10.	The Consumer Protection Act, 2086 and its later versions.

Question Paper Pattern (External only)	
Duration: 3 hrs	**Max**: 75 marks
Section A (5x5=25)	
Short notes Either – Or/ Type - Question from each unit	
Section B(5X10=50)	
Essay type Either – Or/ Type - Question from each unit	

Subject code : 23UCH6Z1

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**PROJECT**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	VI	-	-	5

MARKS DISTRIBUTION	
Project Report	60
Viva-Voce	20
Total	80

Programme Code: 04		B.Sc. Chemistry		
Title of the paper		**JOB ORIENTED COURSE (JOC) – TEXTILE CHEMISTRY**		
Batch		**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024		3	45	

Course Objectives

1.	To know about manufacture and properties of natural fibres (vegetable fibres, animal fibres) and synthetic fibres.
2.	To learn preparatory process before dying.
3.	To know the principles of bleaching and dyeing.

Course Outcomes (CO)

K1 – K5	CO1	Gain the knowledge about both synthetic and natural fibres.
	CO2	To know about Regenerated And Synthetic Fibres.
	CO3	Understand about scouring and desizing.
	CO4	Learn about bleaching.
	CO5	Basic ideas about dyeing

UNIT – I	**Vegetable Fibres And Animal Fibres**	**9 Hours**

Definition –classification of textile fibres- essential and desirable properties of textile fibres-Cotton fibre –Physical and Chemical properties, Jute –Purification; physical and chemical properties of jute, silk and wool.

UNIT-II	**Regenerated And Synthetic Fibres**	**9 Hours**

Rayon –different types of rayon and their sources-manufacture of viscose rayon- physical and chemical properties- acetate rayon –manufacture –properties, enprammonium rayon –manufacture and properties. Manufacture – properties and uses of polyamides- polyester-polypropylene and polyacrylonitrile.

UNIT-III	**Preparatory Process Prior To Dyeing**	**9 Hours**

Scouring: Objective of scouring–Process of caustic scouring on open kier machine with sine diagram, scouring with NaOH and Na_2CO_3 ¬-Precautions to be taken before scouring. Desizing using malt extract-merits and demerits of acid and enzyme desizing Singeing –Impurities present in grey cotton and cotton fabric –objects of singeing –process of singeing on gas singeing machine –precautions to be taken during gas singeing.

UNIT-IV	**Principles of Bleaching**	**9 Hours**

Principles of wetting and mechanism of detergency –synthetic detergents –surface active agents-

bleaching processes –bleaching agents-H_2O_2, NaOCl, bleaching powder and bio-bleaching and their properties-bleaching of cotton, rayon, wool and synthetic fibres.		
UNIT-V	**Principles of Bleaching**	**9 Hours**
Colour and chemical constitution –Chromophore and auxochromes-natural and synthetic dyes-dyes –classification, synthesis of dye shift- congored, bismark brown and erifstal violet, theories of dyeing –effect of temperature and salt on dyeing –dyeing of wool, silk and poly-esters-dyeing of cotton with reactive dyes- fastness properties –washing, light, rubbing and perspiration.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Shenai. V.A., (1991) **Chemistry of Dyes and Principles of Dyeing** Vol. II, Mahajan Publishers, Ahmedabad.	
2.	Gopalakrishnan. R. (2012) **sustainable fibursed textiles**, Textile Fibres SSM, Institute of Textile Technology, Mahajan Publishers, Ahmedabad.	
REFERENCE BOOKS:		
1.	1. Shenai. V.A. (1991) **Textile Fibres** (Vol. I), Mahajan Publishers, Ahmedabad.	
2.	2. Shenai. V.A., (1998) **Technology of Beaching**, Mahajan Publishers, Ahmedabad.	

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	M	S	S	S	M
CO2	H	S	S	S	M
CO3	H	S	S	S	M
CO4	H	S	S	S	H
CO5	S	H	S	S	S

S – Strong **H** – High **M** – Medium **L** – Low

Programme Code: 04		B.Sc., Biotechnology (I Year), Physics (II year), Botany (II Year), Biochemistry (II Year)		
Title of the paper		**ALLIED PAPER – I CHEMISTRY - I**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**I/ III**	4	60	**4**

Course Objectives

1.	To understand the fundamentals of Chemical bonding.
2.	To study various types of organic Reaction.
3.	To study the basic principles of thermodynamics and electrochemistry.

Course Outcomes (CO)

K1 – K5	CO1	Understanding the fundamental aspects of chemical bonding and Interhalogen compounds.
	CO2	To acquire knowledge of types for organic reaction
	CO3	Study on the various concepts in Thermodynamics.
	CO4	Study on the various concepts in Electrochemistry.
	CO5	Acquiring knowledge about Fuel gases and Petroleum.

UNIT – I	**Concepts of Chemical bond**	**12 Hours**

1. Chemical Bonding - Molecular orbital theory- bonding, anti-bonding and non - bonding molecular orbitals - Energy order of MO's - Diamagnetism and Para magnetism - Bond order – Molecular orbital configuration of H_2, N_2, O_2 and F_2.
2. Inter halogen Compounds - Types of Inter halogen Compounds. Preparation, properties, uses and Structures of ICl, BrF_3 and IF_5.
3. Preparation, properties, structure and uses of diborane. .

UNIT-II	**Types of Organic Reactions**	**12 Hours**

1. **Electron Displacement Effects**- Inductive effect – Relative strength of aliphatic acid and alkyl amines. Resonance – Condition for resonance - Consequences of resonance - Hyper conjugation – definition and examples- steric effect.
2. **Aromaticity** – Conditions – Huckel's rule - aromaticity of benzene.
3. **Substitution reactions**- Nitration, halogenation, sulfonation and Fridel alkylation and acylation of benzene.

UNIT-III	**Energetics**	**12 Hours**

Introduction - Scope and limitations - Basic terms - system, surroundings - Types of system, state of system, state variables. Thermodynamic processes - Isothermal, Adiabatic, Isobaric, Isochoric and cyclic processes - Reversible and irreversible processes - Spontaneous process. First law of thermodynamics - Mathematical formulation-limitations. Need for Second law - various statements

of second law. Joule - Thomson effect. Enthalpy - Free energy change.		
UNIT-IV	**Electrochemistry**	**12 Hours**
Introduction - Electrolysis - Conductance of electrolytes - Specific conductance, equivalent conductance, molar conductance - Kohlrausch law - Applications - Determination of degree of dissociation - Conductometric titrations. Buffer Solutions and pH: Buffer solutions- buffer action - Determination of pH of buffer solutions - Buffer solutions in living systems. pH definition - Determination by Colorimetric (indicator) method and electrometric method (Conductometric) only. Principles of electroplating and its uses.		
UNIT-V	**Chemistry of Fuels**	**12 Hours**
1. **Fuel gases** - Qualities of good fuel. Advantages of gaseous fuels over solid and liquid fuels. Short accounts of natural gas, water gas, semi water gas, carburetted water gas, producer gas and oil gas (manufacturing details not required). 2. **Petroleum**-Classification of Petroleum-Refining of crude oil-Cottrell's Process-Removal of Sulphur Compounds-Fractional distillation- Fraction by distillation of crude-Important liquid fuels derived from Petroleum-Gasoline Oil-Kerosene Oil-Diesel Oil.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Arun Bahl and B.S.Bahl, (2012) **Advanced Organic Chemistry**, S. Chand and Co., New Delhi,	
2.	R.D.Madhan- **Modern Inorganic Chemistry**, S.Chand and company, New Delhi.	
3.	Jain and Jain, **Engineering Chemistry**, Dhanpat Rai Publication limited, New Delhi.	
4.	Arun Bahl and B.S.Bahl, Essentials of Physical Chemistry, S. Chand and Co., New Delhi,	
5.	Puri B. R. Sharma L. R., M. S. Pathania, (2013) **Principles of Physical Chemistry**, Vishal Publishing Co., New Delhi.	
REFERENCE BOOKS:		
1.	I.L.Finar, (2009) **Organic Chemistry,**Vol.I and II, Addison-Wesley Longman.	

MAPPING (Physics)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	M	M	M	M
CO2	S	H	M	H	M
CO3	S	M	M	M	M
CO4	S	H	H	H	H
CO5	S	H	H	M	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

MAPPING (Biochemistry)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	M	M	M	M
CO2	S	M	M	H	H
CO3	S	M	M	H	M
CO4	S	H	H	H	H
CO5	S	S	H	M	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

MAPPING (Botany)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	H	M	M	M
CO2	S	H	M	H	H
CO3	S	M	H	H	M
CO4	S	H	H	H	H
CO5	S	S	H	M	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

MAPPING (Biotechnology)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	H	M	H	M
CO2	S	S	H	H	H
CO3	S	H	M	H	M
CO4	S	H	H	M	H
CO5	S	H	H	H	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Programme Code: 04		B.Sc., Biotechnology (I Year), Physics (II year), Botany (II Year), Biochemistry (II Year)		
Title of the paper		**ALLIED PAPER – II CHEMISTRY - II**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	II/ IV	4	60	**4**

Course Objectives

1.	To know the fundamentals of Coordination compounds.
2.	To learn about some natural products, amino acids and proteins.
3.	To study about quantitative and qualitative analysis and synthetic polymer.

Course Outcomes (CO)

K1 – K5	CO1	Understanding the fundamental aspects and applications of coordination Chemistry.
	CO2	Study on the various heterocyclic compounds, carbohydrates and amino acids which include their classification, preparation and properties.
	CO3	To gain knowledge about amino acids and vitamins.
	CO4	To understand theoretical aspects of quantitative and qualitative analysis
	CO5	Acquire the knowledge about synthetic polymers, fibers and plastics

UNIT – I	**Coordination compounds**	**12 Hours**

Complexes (Mononuclear complexes only)

1. General aspects- central metal atom, Ligand- types of ligands. Coordination number of central metal atom, oxidation number of central metal atom – Nomenclature (IUPAC system)
2. Theories of Complexes - Werner's theory, Sidgwick theory - EAN rule, Pauling's theory and its application to Potassium Hexacyanoferrate (II), Tetracarbon nickel (0).
3. Chelation- Definition, examples. EDTA and its applications.
4. Biological role of Hemoglobin and Chlorophyll.

UNIT-II	**Chemistry of Natural Products**	**12 Hours**

1. Heterocyclic compounds – Nomenclature - Preparation - Properties and uses of Furan, Thiophene, Pyrrole and Pyridine. Comparison of the basicities of Pyrrole and Pyridine with amines.
2. Carbohydrates- Classification - Preparation and reactions of glucose and fructose - Interconversion of glucose to fructose and vice versa.

UNIT-III	**Amino acids**	**12 Hours**

1. ***Amino acids** – **Classification** - Preparation - Gabriel Phthalimide synthesis, Strecker synthesis, Amination of α - halo acid-properties.
2. Preparation of Peptides - Bergmann method.

3. Vitamins-Definition-Classification-Sources-Deficiency-Disease.		
UNIT-IV	**Quantitative and Qualitative analysis**	**12 Hours**
1. Quantitative analysis: Various concentration terms - Normality, Molality, Molarity, mole fraction. Volumetric principle, concept of equivalent weight, standard solution, primary standard and secondary standard. **2. Qualitative analysis**: Detection of nitrogen, sulphur, phosphorus and halogens. Identification and properties of functional groups (Amide, Diamide, Carbohydrate, mono and dicarboxylic acid, amine, phenol)		
UNIT-V	**Polymers**	**12 Hours**
1. Synthetic polymers – Nomenclature, Types of polymers – Addition and condensation polymerization. **2. Synthetic fibers** – ***Important requirement** of a fiber, difference between natural and synthetic fiber, properties of synthetic fiber, Preparation of nylon 6 and nylon 6,6. **3. Synthetic plastics** – Classification , thermosetting and thermoplastic plastics – differences, properties		
****self study portion***		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Arun Bahl and B.S.Bahl, (2012) **Advanced Organic Chemistry**, S. Chand and Co., New Delhi,	
2.	R.D.Madhan- **Modern Inorganic Chemistry**, S.Chand and company, New Delhi.	
3.	R. Gopalan, P.S. Subramanian & K. Rengarajan, **Elements of Analytical Chemistry**, Sultan Chand & Sons,	
4.	Arun Bahl and B.S.Bahl, Essentials of Physical Chemistry, S. Chand and Co., New Delhi,	
5.	Puri B. R. Sharma L. R., M. S. Pathania, (2013) **Principles of Physical Chemistry**, Vishal Publishing Co., New Delhi.	
REFERENCE BOOKS:		
1.	I.L.Finar, (2009) **Organic Chemistry,**Vol.I and II, Addison-Wesley Longman.	
2.	R. Gopalan and V. Ramalingam, Concise Coordination Chemistry, Vikas Publishing House Pvt..Ltd, New Delhi.	

MAPPING (Physics)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	H	M	H	M
CO2	S	S	H	H	H
CO3	S	H	M	H	M
CO4	S	H	H	M	H
CO5	S	H	H	M	H
S – Strong	**H** – High	**M** – Medium	**L** – Low		

MAPPING (Biochemistry)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	H	M	H	M
CO2	S	M	H	M	H
CO3	S	H	H	H	M
CO4	S	H	H	M	M
CO5	S	H	H	H	M
S – Strong	**H** – High	**M** – Medium	**L** – Low		

MAPPING (Botany)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	M	M	H	H
CO2	S	S	H	M	H
CO3	H	H	H	H	M
CO4	S	H	H	M	H
CO5	S	H	H	H	M
S – Strong	**H** – High	**M** – Medium	**L** – Low		

MAPPING (Biotechnology)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	H	M	H	M
CO2	S	S	H	H	H
CO3	S	H	M	H	M
CO4	S	H	H	M	H
CO5	S	H	H	M	H
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Subject code : 23UCH2AL/ 23UCH4AL

Programme Code: 04		B.Sc., Biotechnology (I Year), Physics (II year), Botany (II Year), Biochemistry (II Year)		
Title of the paper		**ALLIED PRACTICAL – I VOLUMETRIC AND ORGANIC ANALYSIS**		
Batch	**Semester**	**Hours / Week**	**Total Hours**	**Credits**
2023 - 2024	**II/ IV**	3	90	4

Course Objectives

1.	To demonstrate the basic laboratory technique of titration.
2.	To gain deep knowledge about analysis of organic substances.
3.	To identify the functional groups of unknown compounds.

Course Outcomes (CO)

K1 – K5	CO1	Remember the basics of volumetric titrations.
	CO2	Studying the use of indicators for various titrations.
	CO3	Understanding about preliminary analysis of organic compounds.
	CO4	Identification of the functional groups.
	CO5	Practice for getting accuracy in volumetric estimations

Volumetric Analysis	
1.	Estimation of Sodium hydroxide using standard Sodium Carbonate Solution
2.	Estimation of Hydrochloric acid-standard Oxalic acid
3.	Estimation of Oxalic acid- Standard sulphuric acid.
4.	Estimation of ferrous Sulphate –Standard Mohr salt solution.
5.	Estimation of Oxalic acid- Standard ferrous Sulphate solution.
6.	Estimation of Potassium permanganate- Standard sodium hydroxide solution
Organic Analysis	
1.	Detection of elements (N, S, Halogens).
2.	To distinguish between Aliphatic and Aromatic.
3.	To distinguish between Saturated and unsaturated.
4.	Functional group test for Phenol, acids, (mono and di), aromatic primary amine, amide (mono and di), dextrose.
5.	Systematic analysis of Organic compounds containing one functional group and characterization by confirmatory tests.

TEXT BOOKS :

1. V. Venkateswaran, R. Veeraswmay, A.R. Kulandaivelu,(1997) **Basic Principles of Practical Chemistry**, New Delhi, Sultan Chand and Sons.

REFERENCE BOOKS:

1. N. S. Gnanapragasam and G. Ramamoorthy, (2006) **Organic Chemistry Lab manual,** S. Viswanathan Private Limited, Chennai.

Time: 3 Hours			**Max. Marks: 25**
Distribution of Total Marks 25			
	Record	5	
	Volumetric	10	
	Organic Analysis	10	
Volumetric – 10 Marks			
	Error up to 2%	10	
	3%	8	
	4%	6	
	>4%	4	
Reduce 1 mark for each arithmetic error			
For wrong or no calculation, reduce 25% of the marks awarded. (Here, the examiners have to do the calculation and then, have to award marks)			
Organic Analysis - 10 Marks			
	Preliminary Tests	1	
	Aliphatic / Aromatic	2	
	Saturated / Unsaturated	2	
	Special elements	2	
	Functional group	3	

MAPPING (Physics)

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	H	M	H	M
CO2	S	S	H	H	H
CO3	S	H	M	H	M
CO4	S	H	H	M	H
CO5	S	H	H	H	M

S – Strong **H** – High **M** – Medium **L** – Low

MAPPING (Biochemistry)

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	H	M	H	M
CO2	S	M	H	M	H
CO3	S	H	H	H	M
CO4	S	H	H	M	M
CO5	S	H	H	M	S

S – Strong **H** – High **M** – Medium **L** – Low

MAPPING (Botany)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	M	M	H	H
CO2	S	S	H	M	H
CO3	H	H	H	H	M
CO4	S	H	H	M	H
CO5	S	H	H	H	M
S – Strong	**H** – High		**M** – Medium	**L** – Low	

MAPPING (Biotechnology)					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	H	M	H	M
CO2	S	S	H	H	H
CO3	S	H	M	H	M
CO4	S	H	H	M	H
CO5	S	H	S	S	M
S – Strong	**H** – High		**M** – Medium	**L** – Low	

Kongunadu Arts and Science College, Coimbatore

Title	CERTIFICATE PROGRAMME – DISASTER MANAGEMENT						
Scheme of Curriculum							
Duration: 6 months							
Sub. Code	**Subject**	**Instr. Hours/week**	**Internal**	**External**	**Max marks**	**Exam . Hrs.**	**Credit**
23CDM101	Theory 1 – Disaster Management and Sustainable Development	4	50	50	100	3	2
23CDM102	Theory 2 – Disaster Preparedness and response	4	50	50	100	3	2
23CDM103	Theory 3 – Disaster recovery	4	50	50	100	3	2
				Total	300		6

Question Paper Pattern*
Theory: Ext. Max. Marks: 75 Marks
Section – A (20 x 1 = 20 marks)
MCQ 10/ Fill ups 5/ One ward 5
Section – B (5 x 5 =25 marks) Short Answers (5 Out of 8)
Section – C (3x 10 = 30 marks) Descriptive / Detailed (Three out of 8)
Practical: Ext. Max. Mark: 60 Marks
Experiment: 50 marks
Record: 10 marks
*Question paper pattern as resolved by the Boards of Studies concerned.

Subject code : 23CDM101

Disaster Management		
Title of the paper :	**THEORY 1 – DISASTER MANAGEMENT AND SUSTAINABLE DEVELOPMENT**	
Hours / Week		**Total Hours**
4		60

Course Objectives

1.	To understand the basic aspects of History and Case Studies of Disasters and Pipeline Disasters and oil Spills.
2.	To learn about Climate Changes and Disasters and gain knowledge about Disaster Management Education.
3.	To study about Concept and benefits of Corporate Social Responsibility (CSR).

Course Outcomes (CO)

K1 – K5	CO1	Understand the History and Case Studies of Disasters
	CO2	To understand the Pipeline Disasters and oil Spills & Land degradation and Droughts.
	CO3	Gain the knowledge about Climate Changes and Disasters.
	CO4	Study the basic principles of Disaster Management Education.
	CO5	Explore the Concept and benefits of Corporate Social Responsibility (CSR).

UNIT – I	**12 Hours**

History and Case Studies of Disasters

Introduction, Classification of disasters, History of Disasters, Orissa super Cyclone, Community – based Disaster preparedness (CBDP), Tenth Five Year plan (2002 – 07) and Orissa Super Cyclone,

Floods in Bangladesh

Sustainable Development and Disaster Reduction

Sustainable Development, Sustainable Development Goals(SDGs), Sustainable Development Goals report 2018, World conference on Disaster Risk Reduction , Measures to promote Sustainable Development, International Efforts and Co – operation.

UNIT-II	**12 Hours**

Pipeline Disasters and oil Spills

Pipeline Transportation, Oil Spills, Nigerian Dilemma

Land degradation and Droughts

Land pollution and Degradation, Consequences of Land degradation, Strategies to check land degradation, Droughts: causes, consequences and remedial strategies, programmes for treatment of degraded lands.

UNIT-III	**12 Hours**

Climate Changes and Disasters

Climate Changes and Human Settlements, Climate Change and Water Supply, Occurrence of Droughts, River Hydrology and Inland flooding, Cyclonic storms, Strom Surges and Coastal Flooding, Sea level flooding, India's approach to climate change, National Action plan onclimate change (NAPCC), Eight National Missions of NAPCC.

UNIT-IV	**12 Hours**

Disaster Management Education

Role of education for Disaster Preparedness, Disaster Management education in India, National Disaster Management Training and education for Disaster preparedness, public education, training, rehearsal and experience for disaster preparedness, education, experience and risk perception for disaster preparedness.

UNIT-V	**12 Hours**

Concept and benefits of Corporate Social Responsibility (CSR)

CSR – Definition, main features of CSR, origins of CSR, Approaches to CSR, principle to Triple Bottom Line (TBL), potential business benefits of CSR, Convergence and Transparency of CSR practices, Globalization and CSR, promoting CSR : The challenges.

self study portion

Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class

TEXT BOOKS :

1.	Niranjan Sahoo (2019) **Disaster Management and Sustainable Development**, New Century Publications, New Delhi.
2.	**Disaster Management Guidelines**, GOI-UND Disaster Risk Program (2009-2012)
3.	Damon, P. Copola, (2006), **Introduction to International Disaster Management**, Butterworth Heineman.

REFERENCE BOOKS:

1.	Gupta A.K., Niar S.S and Chatterjee S. (2013) **Disaster management and Risk Reduction**, Role of Environmental Knowledge, Narosa Publishing House, Delhi.
2.	Murthy D.B.N. (2012) **Disaster Management**, Deep and Deep Publication PVT. Ltd. New Delhi. 3.Modh S. (2010) Managing Natural Disasters, Mac Millan publishers India ltd

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	M
CO2	S	S	S	S	S
CO3	S	S	M	H	M
CO4	S	S	H	M	S
CO5	S	S	H	H	S
S – Strong		**H** – High	**M** – Medium	**L** – Low	

Subject code : 23CDM102

Disaster Management			
Title of the paper :	**THEORY 2 – DISASTER PREPAREDNESS AND RESPONSE**		
Hours / Week		**Total Hours**	
4		60	

Course Objectives

1.	To know about the region of Natural disasters and study Safety engineering and analysis techniques.
2.	To have insight about Natural disaster effects and fighting against threats and acquire knowledge about Health care and safety.
3.	To know about National disaster relief strategy and general preparedness.

Course Outcomes (CO)

K1 – K5	CO1	Understand the Types of disasters and causes of disasters.
	CO2	Gain the knowledge about Safety engineering and analysis techniques.
	CO3	Study about the Natural disaster effects and fighting against threats.
	CO4	Know about Health care and safety.
	CO5	Exploring the various National disaster relief strategy and general preparedness.

UNIT – I	**Natural disasters**	**12 Hours**

Types of disasters – causes of disasters – Human made disaster – role of education – Bush fire, earth quake, famine, floods , tsunami, droughts, cyclone, avalanches, volcanic eruption, disaster risk management, real kinship, ritual kinship, non-government organizations, financial resources, environmental issues.

UNIT-II	**Measures for safety**	**12 Hours**

Safety engineering, analysis techniques, safety certification, preventing failure, safety and reliability, containing failure, evolution of safety, evolution of safety management, safety organization in industry, safety culture, function of safety, safety benefits.

UNIT-III	**Natural disaster effects and fighting against threats**	**12 Hours**

Natural disaster effects as Stressors, conceptual guidelines and principles, threat phase, local disaster management cell, current practices, strategic thinking, government response in disasters, types of response.

UNIT-IV	**Health care and safety**	**12 Hours**

Health hazards, physical and mechanical hazards, identifying hazards, conformed space standard, susceptibility of the population, determination of the health service, climatic events, levels of intervention, curative actions, health education.

UNIT-V	**National strategy for relief**	**12 Hours**
National disaster relief strategy and general preparedness, general principles, disasters alert, relief phase, myths and reality, key aspects effective response, medical health care, food supplies, water supplies, hygiene, environmental sanitation, construction equipment, communications, logistics systems and facilities.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	K.K. David (2020), **Disaster Management and Preparedness**, Rajat publicatios, New Delhi.	
2.	**Disaster Management Guidelines**, GOI-UND Disaster Risk Program (2009-2012)	
3.	Damon, P. Copola, (2006) **Introduction to International Disaster Management**, Butterworth Heineman.	
REFERENCE BOOKS:		
1.	Gupta A.K., Niar S.S and Chatterjee S. (2013) **Disaster management and Risk Reduction**, Role of Environmental Knowledge, Narosa Publishing House, Delhi.	
2.	Murthy D.B.N. (2012) **Disaster Management**, Deep and Deep Publication PVT. Ltd. New Delhi. 3.Modh S. (2010) Managing Natural Disasters, Mac Millan publishers India LT	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	M
CO2	S	S	S	S	S
CO3	S	S	M	H	M
CO4	S	S	H	M	S
CO5	S	S	H	H	S
S – Strong	**H** – High		**M** – Medium	**L** – Low	

Sub `ject code : 23CDM103

Disaster Management		
Title of the paper :	**THEORY 3 – DISASTER RECOVERY**	
Hours / Week	**Total Hours**	
4	60	

Course Objectives

1.	To learn group Causes of disaster and study about Disaster recovery plan.
2.	To know about Role of technology in disaster recovery management and study about Environmental disaster management.
3.	To learn about Disaster management to psychological perspectives.

Course Outcomes (CO)

K1 – K5	CO1	Gain the knowledge about Disaster recovery.
	CO2	Understand the basic aspects of Disaster recovery plan.
	CO3	Analyze and apply Role of technology in disaster recovery management.
	CO4	Understand about Brief history of the environment movement.
	CO5	To meet the contemporary challenges on Disaster management to psychological perspectives.

UNIT – I	**Introduction to disaster recovery**	**12 Hours**

Causes of disaster, Introduction, self-reliance, UN – DHA objective, operational fire management systems and organizations, role of military and civil defense assets, importance of disaster recovery planning, control measures, strategies, recovery time objective, recovery time actual, recovery consistency objective, recovery point objectives.

UNIT-II	**Disaster recovery plan**	**12 Hours**

Objectives, relationship to the business continuity plan, benefits, types of plan, types of disasters, planning methodology, caveats and controversies, disaster management apparatus, Sahana Foss disaster management system.

UNIT-III	**Role of technology in disaster recovery management**	**12 Hours**

IT network disaster recovery, emergency management systems, GPS, applications of GPS to disaster management, thermal infrared remote sensing technology, population dynamics model for disaster management. Integral view of space technologies.

UNIT-IV	**Environmental disaster management**	**12 Hours**

Brief history of the environment movement, environmental and the four phases of emergency management, environmental in the disaster context, the news of environmental management, development and disaster risk.

UNIT-V	**Disaster management to psychological perspectives**	**12 Hours**
Event of psychological impact of disasters, Psychological interventions to victims, Forms of psycho pathology resulting from disaster, psychological perspectives and disasters, disaster perspective and disaster relief.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		

TEXT BOOKS :

1.	Sachin Sehdev Paarikh (2018), Disaster Recovery, Random Publications, New Delhi
2.	Disaster Management Guidelines, GOI-UND Disaster Risk Program (2009-2012)
3.	Damon, P. Copola, (2006) Introduction to International Disaster Management, Butterworth Heineman.

REFERENCE BOOKS:

1.	Gupta A.K., Niar S.S and Chatterjee S. (2013) **Disaster management and Risk Reduction**, Role of Environmental Knowledge, Narosa Publishing House, Delhi.
2.	Murthy D.B.N. (2012) **Disaster Management**, Deep and Deep Publication PVT. Ltd. New Delhi. 3.Modh S. (2010) Managing Natural Disasters, Mac Millan publishers India LT

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	S	H	M	M
CO2	S	S	S	S	S
CO3	S	S	M	H	M
CO4	S	S	H	M	S
CO5	S	S	H	H	S

S – Strong **H** – High **M** – Medium **L** – Low

Kongunadu Arts and Science College, Coimbatore

Title :	**CERTIFICATE PROGRAMME – INSTRUMENTAL METHODS OF CHEMICAL ANALYSIS**

Scheme of Curriculum

Duration: 6 months

Sub. Code	**Subject**	**Instr. Hours/week**	**Internal**	**External**	**Max marks**	**Exam. Hrs.**	**Credits**
23CIM101	Theory 1– Analytical Chemistry	4	50	50	100	3	2
23CIM102	Theory 2 – Spectroscopic techniques	4	50	50	100	3	2
23CIM1CL	Practical 1 Instrument and Chemical methods in day to day activity	4	50	50	100	3	2
				Total	300		6

Question Paper Pattern*
Theory: Ext. Max. Marks: 75 Marks
Section – A (20 x 1 = 20 marks)
MCQ 10/ Fill ups 5/ One ward 5
Section – B (5 x 5 =25 marks) Short Answers (5 Out of 8)
Section – C (3x 10 = 30 marks) Descriptive / Detailed (Three out of 8)
Practical: Ext. Max. Mark: 60 Marks
Experiment: 50 marks
Record: 10 marks
*Question paper pattern as resolved by the Boards of Studies concerned.

Subject code : 23CIM101

Instrumental Methods Of Chemical Analysis			
Title of the paper :	**THEORY 1 – ANALYTICAL CHEMISTRY**		
Hours / Week		**Total Hours**	
4		60	

Course Objectives

1.	To understand the key features of Analytical chemistry and know the basics of Errors, Accuracy and Precision.
2.	To identify different Separation techniques and describe about important purification techniques.
3.	To understand important of Analytical biochemistry and industrial process.

Course Outcomes (CO)

K1 – K5	CO1	Understand the various types of analytical methods.
	CO2	To know the Knowledge about basics of Errors, Accuracy and Precision.
	CO3	Analyze the importance of Separation techniques.
	CO4	Describe about various types of purification techniques.
	CO5	To gain knowledge about Analytical biochemistry and industrial process

UNIT – I	**Introduction to Analytical Chemistry**	**12 Hours**

Introduction – types of analytical methods, instruments for analysis, uncertainty in instrumental measurements, senility and detection limits for instruments.

UNIT-II	**Errors, Accuracy and precision**	**12 Hours**

Errors – types of errors, significant figures, precision and accuracy, methods of expressing accuracy, methods of expressing precision, confidence limit, photometric errors.

UNIT-III	**Separation techniques**	**12 Hours**

Precipitation, solvent extraction, dessicant, types of dessicant, relative efficiencies of dessicant, trying paper and temperature, choice of dissicants, distillation, theory of distillation, recrystallization, sublimation.

UNIT-IV	**Purification techniques**	**12 Hours**

Principles and techniques of semi micro analysis, filtration of precipitates, washing of precipitates, heating solutions, evaporation, transferring residue, methods of precipitating sulphides.

UNIT-V	**Analytical biochemistry and industrial process**	**12 Hours**
Definition of pH, pH in biological system, buffer system in animal kingdom, redox titration, electroplating-principle and process of electroplating, applications of electroplating.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Vogel, Arthur I: A **Text book of Quantitative Inorganic Analysis** (Rev. by G.H. Jeffery and others) 5th Ed. The English Language Book Society of Longman.	
2.	Willard, Hobert H. et al.: **Instrumental Methods of Analysis**, 7th Ed. Wardsworth Publishing Company, Belmont, California, USA, 1988.	
3.	Christian, Gary D; **Analytical Chemistry**, 6th Ed. John Wiley & Sons, New York, 2004.	
REFERENCE BOOKS:		
1.	Harris, Daniel C: **Exploring Chemical Analysis**, Ed. New York, W.H. Freeman, 2001.	
2.	2. Khopkar, S.M. **Basic Concepts of Analytical Chemistry**. New Age, International Publisher, 2009.	
3.	Skoog, D.A. Holler F.J. and Nieman, T.A. **Principles of Instrumental Analysis**, Thomson Asia Pvt. Ltd. Singapore.	
4.	Mikes, O. & Chalmes, R.A. **Laboratory Hand Book of Chromatographic & Allied Methods**, Elles Harwood Ltd. London.	
5.	Ditts, R.V. **Analytical Chemistry** – Methods of separation.	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	H
CO2	S	M	S	M	S
CO3	S	M	S	M	S
CO4	S	S	S	S	S
CO5	S	S	H	S	S
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Subject code : 23CIM102

Instrumental methods of chemical analysis			
Title of the paper :	**THEORY 2 – SPECTROSCOPY AND CHROMATOGRAPHIC TECHNIQUES**		
Hours / Week		**Total Hours**	
4		60	

Course Objectives

1.	To study Ultra – Violet and visible spectroscopy and study about Infrared spectroscopy.
2.	To understand about Nuclear Magnetic Resonance (NMR) spectroscopy.
3.	To study about various types chromatography and understand experimental techniques of column chromatography,

Course Outcomes (CO)

K1 – K5	CO1	Understanding the fundamental aspect Ultra – violet and visible spectroscopy.
	CO2	Learn about Fundamental concepts of Infrared spectroscopy
	CO3	Acquire the knowledge of Nuclear Magnetic Resonance (NMR) spectroscopy.
	CO4	Study on the various types of chromatography.
	CO5	To inculcate knowledge about Column chromatography.

UNIT – I	**Ultra – violet and visible spectroscopy**	**12 Hours**

Electromagnetic spectrum and absorption of radiations, Principle of ultraviolet spectroscopy, the absorption laws- Lambert's law and Beer's law. Selection rules, instrumentation – Block Diagram, theory of electronic spectroscopy, types of electronic transitions.

UNIT-II	**Infrared spectroscopy**	**12 Hours**

Fundamental concepts of Infrared spectroscopy, molecular vibrations, vibrational frequency, number of fundamental vibrations, selection rules, Scanning of infrared spectrum (instrumentation), finger print region.

UNIT-III	**Nuclear Magnetic Resonance (NMR) spectroscopy**	**12 Hours**

Introduction, conditions of resonance, Solvents used in NMR, relaxation process – spin –spin relaxation, spin – lattice relaxation and quadrupole relaxation. Number of signals, instrumentation, splitting of the signals in pure ethanol and chloroethane

UNIT-IV	**Thin layer chromatography**	**12 Hours**

Introduction, definition, types, thin layer chromatography – experimental techniques - coating materials - preparation of thin layer in plates – activation of adsorbents – purification of silica gel – sample application – development tank – solvent systems – plate development – detection of components, evaluation of chromatography.

UNIT-V	**Column chromatography**	**12 Hours**
Column chromatography – principle, experimental techniques – apparatus – adsorbents – preparation of adsorption columns – solvents used in successive elution – gradient elution.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		
TEXT BOOKS :		
1.	Y.R, Sharma, (2007) **Elementary Organic Spectroscopy**	
2.	Jag Mohan, (2013) **Organic Spectroscopy – Principles and Applications**, Narosa publishing house.	
3.	Gurdeep R. Chatwal, (2002) **Instrumental Methods of Chemical Analysis**, Himalaya publishing house, Delhi.	
REFERENCE BOOKS:		
1.	P.S. Sindhu, Elements of (2007) **Molecular Spectroscopy**, New Age International Publishers, New Delhi.	
2.	H.S. Randhana, (2003) **Modern Molecular Spectroscopy**, Macmillan India Ltd, New Delhi.	

MAPPING					
PSO / CO	**PSO 1**	**PSO 2**	**PSO 3**	**PSO 4**	**PSO 5**
CO1	S	S	H	M	H
CO2	S	M	S	M	S
CO3	S	M	S	M	S
CO4	S	S	S	S	S
CO5	S	S	H	S	S
S – Strong	**H** – High	**M** – Medium	**L** – Low		

Subject code : 23CIM1CL

Instrumental methods of chemical analysis	
Title of the paper :	**PRACTICAL - 1 INSTRUMENT AND CHEMICAL METHODS IN DAY TO DAY ACTIVITY**

Hours / Week	**Total Hours**
4	60

Course Objectives

1.	Transformation of theoretical knowledge gain to practical aspects and have experience in handling organic compounds.
2.	To determine the dissolved oxygen in different types of water.
3.	To know about fat content in milk using Lactometer and analysis techniques and understand about Estimation of Hardness of water.

Course Outcomes (CO)

K1 – K5	CO1	Understand the basic principles of Instrumental Methods of Chemical Analysis.
	CO2	Understand about the Colorimetric experiments.
	CO3	Determination of boiling point of organic compounds.
	CO4	Improve the accuracy of analysis.
	CO5	To gain knowledge about Biological Oxygen Demand (BOD).

Experiments

1.	Detection of boiling point of organic compounds
2.	Detection of melting point of Inorganic and Organic compounds
3.	Colorimetric experiments
4.	Estimation of dissolved oxygen in different types of water
5.	Detection of fat content in milk using Lactometer and analysis techniques
6.	Detection of Biological Oxygen Demand (BOD)
7.	Detection of Hardness of water

UNIT-I	**Electromagnetic spectrum**	**12 Hours**

Electromagnetic spectrum and absorption of radiations, Principle of ultraviolet spectroscopy, the absorption laws- Lambert's law and Beer's law. Selection rules, instrumentation – Block Diagram, theory of electronic spectroscopy, types of electronic transitions.

UNIT-II	**Infrared spectroscopy**	**12 Hours**

Fundamental concepts of Infrared spectroscopy, molecular vibrations, vibrational frequency, number of fundamental vibrations, selection rules, Scanning of infrared spectrum (instrumentation), finger print region.

UNIT-III	**Nuclear Magnetic Resonance (NMR) spectroscopy**	**12 Hours**

Introduction, conditions of resonance, Solvents used in NMR, relaxation process – spin –spin relaxation, spin – lattice relaxation and quadrupole relaxation. Number of signals, instrumentation, splitting of the signals in pure ethanol and chloroethane

UNIT-IV	**Thin layer chromatography**	**12 Hours**
Introduction, definition, types, thin layer chromatography – experimental techniques - coating materials - preparation of thin layer in plates – activation of adsorbents – purification of silica gel – sample application – development tank – solvent systems – plate development – detection of components, evaluation of chromatography.		
UNIT-V	**Column chromatography**	**12 Hours**
Column chromatography – principle, experimental techniques – apparatus – adsorbents – preparation of adsorption columns – solvents used in successive elution – gradient elution.		
self study portion		
Teaching Methods : Smart Class Room/ Power point Presentation/ Seminar/ Quiz/ Discussion/ Flipped Class		

TEXT BOOKS :

1.	Vogel, A.I. **Quantitative Organic Analysis,** Part 3, Pearson (2012).
2.	Mann, F.G. & Saunders, B.C. **Practical Organic Chemistry**, Pearson Education (2009
3.	Ahluwalia, V.K. & Aggarwal, R. **Comprehensive Practical Organic Chemistry**: Preparation and Quantitative Analysis, University Press (2000).
4.	Furniss, B.S.; Hannaford, A.J.; Smith, P.W.G.; Tatchell, A.R. **Practical Organic Chemistry**, 5th Ed., Pearson (2012)

REFERENCE BOOKS:

1.	Ahluwalia, V.K. & Dhingra, S. Comprehensive Practical Organic Chemistry: Qualitative Analysis, University Press (2000).
2.	H.S. Randhana, (2003) **Modern Molecular Spectroscopy**, Macmillan India Ltd, New Delhi.

Time: 3 Hours	**Max. Marks: 50**
Distribution of Total Marks - 50	
Record	**05 mark**
Practical	**45 mark**

MAPPING

PSO / CO	PSO 1	PSO 2	PSO 3	PSO 4	PSO 5
CO1	S	S	H	M	H
CO2	S	S	M	H	H
CO3	S	S	H	M	H
CO4	S	S	M	H	H
CO5	S	S	H	H	H

S – Strong **H** – High **M** – Medium **L** – Low